

Chapter 10 Chemical Quantities Review

Part 1: Vocabulary

Label the following descriptions with the correct vocabulary word.

- | | |
|---|---|
| _____ mass | _____ percent composition. |
| _____ empirical formula | _____ empirical formula |
| _____ molar mass | _____ standard temperature and pressure (STP, 1 atm, 0°C) |
| _____ 10 ²⁴ | |
| _____ is a description of the relative amounts of each element in a compound _____ percent composition | |
| _____ the lowest whole number ratio of the atoms of the elements in a compound _____ empirical formula | |
| _____ the volume of 1 mole of any gas at STP _____ 22.4 L | |
| _____ the smallest particle of a compound is a compound _____ molar mass | |
| _____ the number representing 6.02 $\times 10^{23}$ representative particles of a substance _____ Avogadro's number | |
| _____ 1 mole = 6.02 $\times 10^{23}$ particles _____ empirical formula | |
| _____ the number representing the relative mass of an element or compound (STP, 1 atm) _____ standard temperature and pressure (STP, 1 atm) | |

Part 2: Section Concept Review

10.1 The Mole: A Measurement

Completion Use the completion strategy to strengthen knowledge of the terms and your understanding of the concepts involved with the terms of the definitions completed with a term (don't guess) or explain.

Measuring whole units of something, or mass, and its volume for a single quantity called the _____ mole. _____ Representative representative particles like large or a substance is _____ Avogadro's number _____

It is the mass of _____ of a compound, substance and together the _____ percent mass _____ of the substance of the compound. Other use techniques the substance the entire (the whole) the _____ molar mass _____ of the compound. These are _____ 6.02 $\times 10^{23}$ _____ representative particles the mass of one substance.

True or False

_____ 1 mole of any substance is 6.02 $\times 10^{23}$ particles.

_____ 1 mole of substance contains 6.02 $\times 10^{23}$ atoms.

_____ 1. The representative particles of a substance is the molecule. **Incorrect**

_____ 1. A mole of 100 is composed of one mole of carbon and four atoms of oxygen.

_____ 1. The molar mass of nitrogen gas (N₂) is **Incorrect** **nitrogen's atomic mass**

Problems

Chemical Quantities Ch Vocab Review Key

CO Houle



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