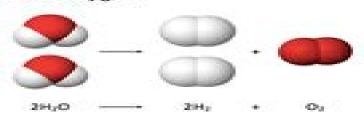
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Name: _____

In chemistry, we learn about how atoms form bonds with one another in order to satisfy their need for complete electron shells. For example, oxygen's need for 2 additional electrons is filled by the bond formed with two hydrogen atoms, both having one "extra" electron. The bond between these three atoms creates the molecule we know as water. You've been around water your entire life, but have you ever seen a water molecule split back into hydrogen and oxygen?

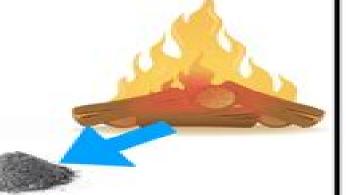
Most likely you haven't! That's because these chemical bonds are not easily broken. Splitting hydrogen and oxygen isn't like picking the pineapple off your pizza. When molecules are broken apart or built up from separate pieces, it's known as a **chemical reaction**. Even if you've never witnessed the splitting of water, chemical reactions happen around you all the time!



When a chemical reaction has occurred, the starting substances (reactants) are modified and become something different (products). Signs that a chemical reaction have taken place are: temperature increase or decrease, gas production (bubbles), color change, odor, & formation of a precipitate (solids).

Let's try to categorize some examples. Write YES if you think a chemical reaction has occurred, NO if it hasn't.

- 1. Cutting a piece of paper
- 2. Burning a log
- 3. Rust forming from iron
- Adding sugar to your tea _____
- Baking a cake
- A banana rotting



Chemical reactions involve breaking apart molecules and combining atoms in new ways to form new materials. One thing must always be true of any chemical reaction: The amount of matter that existed at the start of the reaction must equal the amount of matter that exists at the end. This is known as the **Law of Conservation of Mass**. Sound simple? It seems so, but why when you burn a log do you end up with such a small pile of ash? Where did the rest of that matter go? Think about your prediction before you flip the page.

Chemical Reaction Comprehension

C Cleary

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