

Chapter 8 Chemical Reactions Chapter Review Answers

1. The evolution of energy, the production of a gas, the formation of a precipitate, and a color change.
2. A correctly written chemical equation must represent the known facts, contain the correct formulas for the reactants and products, and satisfy the law of conservation of mass.
3. a. It is a number that appears in front of a formula in a chemical equation.
b. The number of atoms of each type in the formula is multiplied by the coefficient.
4. methane + oxygen $\rightarrow$ carbon dioxide + water;
 $\text{CH}_4(\text{g}) + \text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g})$;
 $\text{CH}_4(\text{g}) + 2\text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + 2\text{H}_2\text{O}(\text{g})$
5. A chemical equation provides the number of molecules, moles, and atoms of each element or compound composing the reactants and products as well as the mole ratios of the reactants to each other and to the products.
6. Word and formula equations have only qualitative meaning; they do not express the relative quantities of a reaction's reactants or products.
7. a. a solution in water
b. a substance that accelerates a chemical reaction but can be recovered unchanged
c. a chemical reaction in which the products re-form the original reactants
8. a. KOH d. CCl_4
b. $\text{Ca}(\text{NO}_3)_2$ e. MgBr_2
c. Na_2CO_3
9. (1) Balance different types of atoms one at a time; (2) balance types of atoms that appear only once on each side of the equation first; (3) balance as single units any polyatomic ions that appear on both sides of the equation; and (4) balance H atoms and O atoms last.
10. a. 6N
b. 4H, 2O
c. 4H, 4N, 12O
d. 2Ca, 4O, 4H
e. 3Ba, 6Cl, 18O
f. 5Fe, 10N, 30O
g. 12Mg, 8P, 32O
h. 4N, 16H, 28, 8O
i. 12Al, 18Se, 72O
j. 12C, 32H
11. a. $2\text{ZnS}(\text{s}) + 3\text{O}_2(\text{g}) \rightarrow 2\text{ZnO}(\text{s}) + 2\text{SO}_2(\text{g})$
b. $2\text{HCl}(\text{aq}) + \text{Ba}(\text{OH})_2(\text{aq}) \rightarrow \text{BaCl}_2(\text{aq}) + 2\text{H}_2\text{O}(\text{l})$
c. $2\text{HNO}_3(\text{aq}) + \text{Ca}(\text{OH})_2(\text{aq}) \rightarrow \text{Ca}(\text{NO}_3)_2(\text{aq}) + 2\text{H}_2\text{O}(\text{l})$
12. a. Solid zinc sulfide reacts with oxygen gas to form solid zinc oxide and sulfur dioxide gas.
b. When solid calcium hydride is added to water, aqueous calcium hydroxide and hydrogen gas are formed.
c. Aqueous silver nitrate mixed with aqueous potassium iodide produces a precipitate of silver iodide and aqueous potassium nitrate.
13. a. $\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$
b. $2\text{Al} + \text{Fe}_2\text{O}_3 \rightarrow \text{Al}_2\text{O}_3 + 2\text{Fe}$
c. $\text{Pb}(\text{CH}_3\text{COO})_2 + \text{H}_2\text{S} \rightarrow \text{PbS} + 2\text{CH}_3\text{COOH}$
14. a. Li_2O_2 is an incorrect formula; $4\text{Li} + \text{O}_2 \rightarrow 2\text{Li}_2\text{O}$
b. H_2Cl_2 is an incorrect formula; $\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$
c. MgO_2 is an incorrect formula, and the equation as written is not balanced; $\text{MgCO}_3 \rightarrow \text{MgO} + \text{CO}_2$
d. I is an incorrect formula for iodine, and the equation is not balanced; $2\text{NaI} + \text{Cl}_2 \rightarrow 2\text{NaCl} + \text{I}_2$
15. a. $4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$
b. $\text{P}_4\text{O}_{10} + 6\text{H}_2\text{O} \rightarrow 4\text{H}_3\text{PO}_4$
c. $\text{Fe}_2\text{O}_3 + 3\text{CO} \rightarrow 2\text{Fe} + 3\text{CO}_2$
16. $\text{CH}_4(\text{g}) + 4\text{Cl}_2(\text{g}) \rightarrow \text{CCl}_4(\text{l}) + 4\text{HCl}(\text{g})$
17. a. $2\text{Mg} + \text{O}_2 \rightarrow 2\text{MgO}$
b. $4\text{Fe} + 3\text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3$
c. $2\text{Li} + \text{Cl}_2 \rightarrow 2\text{LiCl}$
d. $\text{Ca} + \text{I}_2 \rightarrow \text{CaI}_2$
18. (1) In a synthesis reaction, two or more substances combine to form a new substance: $\text{A} + \text{X} \rightarrow \text{AX}$.
(2) In a decomposition reaction, a single compound undergoes a reaction that produces two or more simpler substances: $\text{AX} \rightarrow \text{A} + \text{X}$.
(3) In a single-displacement reaction, one element replaces another element in a compound:
 $\text{A} + \text{BX} \rightarrow \text{AX} + \text{B}$ or $\text{Y} + \text{BX} \rightarrow \text{BY} + \text{X}$

Chemical Reactions Chapter Review Answers

Stanley E. Manahan



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Table of Contents Chemical Reactions Chapter Review Answers

1. Understanding the eBook Chemical Reactions Chapter Review Answers
 - The Rise of Digital Reading Chemical Reactions Chapter Review Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemical Reactions Chapter Review Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemical Reactions Chapter Review Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemical Reactions Chapter Review Answers
 - Personalized Recommendations
 - Chemical Reactions Chapter Review Answers User Reviews and Ratings
 - Chemical Reactions Chapter Review Answers and Bestseller Lists
5. Accessing Chemical Reactions Chapter Review Answers Free and Paid eBooks

- Chemical Reactions Chapter Review Answers Public Domain eBooks
 - Chemical Reactions Chapter Review Answers eBook Subscription Services
 - Chemical Reactions Chapter Review Answers Budget-Friendly Options
6. Navigating Chemical Reactions Chapter Review Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemical Reactions Chapter Review Answers Compatibility with Devices
 - Chemical Reactions Chapter Review Answers Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemical Reactions Chapter Review Answers
 - Highlighting and Note-Taking Chemical Reactions Chapter Review Answers
 - Interactive Elements Chemical Reactions Chapter Review Answers
 8. Staying Engaged with Chemical Reactions Chapter Review Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemical Reactions Chapter Review Answers
 9. Balancing eBooks and Physical Books Chemical Reactions Chapter Review Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemical Reactions Chapter Review Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Chemical Reactions Chapter Review Answers
 - Setting Reading Goals Chemical Reactions Chapter Review Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Chemical Reactions Chapter Review Answers
 - Fact-Checking eBook Content of Chemical Reactions Chapter Review Answers
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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