

Chemical Reactions

Section 9.1 Reactions and Equations

In your textbook, read about evidence of chemical reactions.

For each statement, write *yes* if evidence of a chemical reaction is present. Write *no* if there is no evidence of a chemical reaction.

- | | |
|---------------|---|
| <u>yes</u> | 1. A tomato smells rotten. |
| <u>no</u> | 2. A drinking glass breaks into smaller pieces. |
| <u>no</u> | 3. A piece of ice melts. |
| <u>yes</u> | 4. Drain cleaner is mixed with water and the solution becomes warm. |
| <u>yes</u> | 5. Candle wax burns. |
| <u>yes/no</u> | 6. Molten candle wax solidifies. |
| <u>yes</u> | 7. Green leaves turn yellow and red as the seasons change. |
| <u>yes</u> | 8. Baking powder produces a gas that makes a cake rise. |

In your textbook, read about how to represent chemical reactions and how to balance chemical equations.

Use the terms below to complete the passage. Each term may be used once, more than once, or not at all.

arrow	plus sign	(s)	(l)
reactant	product	(g)	(aq)

The fuel for the space shuttle is hydrogen, which burns in oxygen to produce water vapor and energy. In this chemical reaction, hydrogen is a(n) **(9)** reactant, oxygen is a(n) **(10)** reactant, and water vapor is a(n) **(11)** product. In a chemical equation for this reaction, a(n) **(12)** arrow is used to separate hydrogen and oxygen from water vapor and energy. A(n) **(13)** plus sign is used to separate the symbols for hydrogen and oxygen. A(n) **(14)** (g) symbol is used to tell the state of hydrogen in the reaction, a(n) **(15)** (g) symbol is used for the state of oxygen, and a(n) **(16)** (g) symbol is used for the state of water vapor.

Chemical Reactions Discussion Guide

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Readers' Guide to Periodical Literature Anna Lorraine Guthrie, Bertha Tannehill, Neltje Marie Tannehill Shimer, 1925
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Atkins' Physical Chemistry Peter Atkins, Julio de Paula, 2010 This volume features a greater emphasis on the molecular view of physical chemistry and a move away from classical thermodynamics It offers greater explanation and support in

mathematics which remains an intrinsic part of physical chemistry *Reaction Guide for Organic Chemistry* Michael J. Millam, 1989 **Chemical Reactions and Their Equations** Ingo Waldemar Dagobert Hackh, 1921 *Study Guide [to Accompany] General Chemistry* James E. Brady, 1982 **A Study Guide to Basic Principles of Organic Chemistry** John D. Roberts, Marjorie C. Caserio, 1977 **An Introductory Guide to EC Competition Law and Practice** Valentine Korah, 1994 A Study Guide to Organic Chemistry John D. Roberts, Marjorie C. Caserio, 1971 Organic Chemistry Study Guide and Solutions Manual Marc Loudon, Jim Parise, Lisa Bonner, 2024-08-09 The Study Guide and Solutions Manual offers several study tools for students Study Guide Links are brief supplements to text material and provide hints for study shortcuts approaches to problem solving or additional explanations of topics that typically cause difficulty Further Explorations are short discussions that provide additional depth on content in the text Reaction Reviews present the reactions in the chapter in detailed outline form This summary consolidates the reactions of the chapter in one place and presents not only the reactions themselves but also the essential features of their mechanisms and stereochemistry for cases in which these issues are discussed in the text Solutions are provided for all of the over 1600 problems in the text Many include worked out examples in which not only the answer is given but also the detailed approach to the solution is discussed
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bilmeniz gerekenleri bu içeriğimizde yer verdik

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