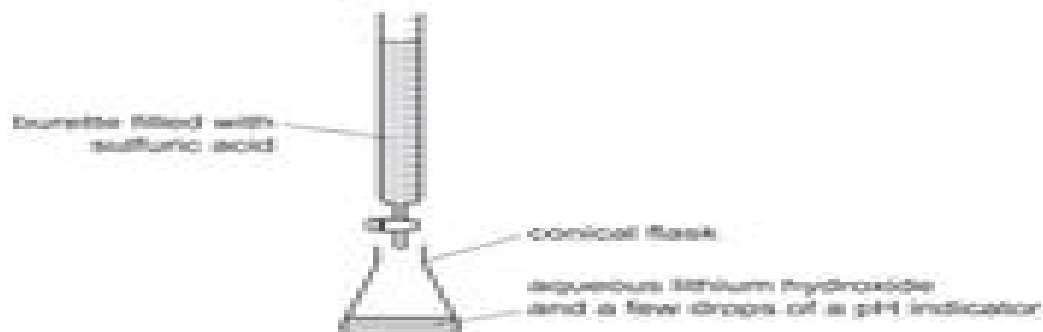


- 7 The soluble salt hydrated lithium sulfate is made by titration from the soluble base lithium hydroxide.



- (a) The sulfuric acid is added slowly from the burette until the indicator just changes colour. The volume of sulfuric acid needed to just neutralise the lithium hydroxide is noted. Describe how you would continue the experiment to obtain pure dry crystals of hydrated lithium sulfate.

[5]

- (b) Using 25.0 cm³ of aqueous lithium hydroxide, concentration 2.48 mol/dm³, 2.20 g of hydrated lithium sulfate was obtained. Calculate the percentage yield, giving your answer to one decimal place.



Number of moles of LiOH used = _____

Number of moles of Li₂SO₄·H₂O which could be formed = _____

Mass of one mole of Li₂SO₄·H₂O = 128 g

Maximum yield of Li₂SO₄·H₂O = _____ g

Percentage yield = _____ %

[4]

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