

CHEM 120 FINAL EXAM

6. (TCO 6) A gas at a temperature of 95 degrees C occupies a volume of 165 mL. Assuming constant pressure, determine the volume at 25 degrees C. Show your work. (Points : 5)

Using Charles' Law, $(V_1/T_1) = (V_2/T_2)$.
First, convert temperature to KELVIN ($T_1 = t_1 + 273$)
Thus, $T_1 = 95 + 273 = 368$.

We have V_1 (165 mL) & $T_2 = (25 + 273) = 298$.
 $V_2 = (V_1 * T_2) / T_1 = (165 \text{ mL} * 298) / 368 = 133.6 \text{ mL}$.

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7. (TCO 6) A sample of helium gas occupies 1021 mL at 719 mmHg. For a gas sample at constant temperature, determine the volume of helium at 745 mmHg. Show your work. (Points : 5)

$1021 \text{ mL} * 719 \text{ mm} / 745 \text{ mmHg} = 985.36 \text{ mL} \approx 985 \text{ mL}$

Using Boyle's law, $P_1 V_1 = P_2 V_2$. We have V_1 (1021 mL), P_1 (719 mmHg) and P_2 (745 mmHg).

0	1644446459	Short	19
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8. (TCO 12) If one strand of a DNA double helix has the sequence T T A G C G A C G C, what is the sequence of the other DNA strand? (Points : 10)

A A T C G C T G C G

1. (TCO 8) 35.0 mL of 0.25 M NaOH is neutralized by 23.6 mL of an HCl solution. The molarity of the HCl solution is (show your work): (Points : 5)

molarity = moles solute / liters solution 0.25 M = moles NaOH / 0.035 L moles NaOH = 0.00875 moles NaOH		▲ ▼
◀		▶

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2. (TCO 1) How many mL are in 3.5 pints? Show your work. (Points : 5)

3.5 pints is equivalent to 1656.116		▲ ▼
1 pint = 473.176 ml 3.5 pints * 473.176 mL = 1656.116 mL		▲ ▼
◀		▶

0	1644446452	Short	6
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