


A balanced chemical equation is a chemist's recipe.

Chemical Equations					
	$\text{N}_2(g)$	+	$3\text{H}_2(g)$	$\longrightarrow$	$2\text{NH}_3(g)$
"Microscopic recipe"	1 molecule $\text{N}_2$	+	3 molecules $\text{H}_2$	$\longrightarrow$	2 molecules $\text{NH}_3$
"Macroscopic recipe"	1 mol $\text{N}_2$	+	3 mol $\text{H}_2$	$\longrightarrow$	2 mol $\text{NH}_3$

- What would happen if two molecules (moles) of  $\text{N}_2$  reacted with three molecules (moles) of  $\text{H}_2$ ?

Experimental Conditions		
	Reactants	Products
Before reaction	 2 molecules $\text{N}_2$  3 molecules $\text{H}_2$	0 molecules $\text{NH}_3$

# Chemistry 12 3 Limiting Reagent And Percent Yield Answers

**Deborah Wiegand**



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