

7. What is the oxidation state of the bold element in each compound below?

a) **Fe**Cl<sub>2</sub> \_\_\_\_\_

b) Na<sub>2</sub>**S**O<sub>3</sub> \_\_\_\_\_

c) H<sub>2</sub>**O**<sub>2</sub> \_\_\_\_\_

d) Na<sub>3</sub>**P**O<sub>4</sub> \_\_\_\_\_

8. Place the following diatomic molecules in order of increasing polarity: CN, CF, O<sub>2</sub>

9. Draw a plot of pressure vs. moles of gas for an ideal gas (label the axes, including units).

10. Circle the correct answer for each of the following:

a) the lowest electronegativity: F, Cl, **P**

b) would be most likely to form a stable anion with a -2 charge: Se, Cl, Mg

c) is diamagnetic: F, Fe<sup>3+</sup>, Na<sup>+</sup>

13. Like all carbonates, barium carbonate can be decomposed to the metal oxide and CO<sub>2</sub>.



If the  $\Delta H_{\text{rxn}}$  of barium carbonate decomposition is 269.3 kJ/mol, how many kJ are required to decompose 10.0 g of BaCO<sub>3</sub>?

# Chemistry 1 Semester 2 Practice Final Exam

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Kentucky**



## Chemistry 1 Semester 2 Practice Final Exam:

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