

## AP Chemistry - Solution Stoichiometry - Problem Set 3

### Acids / Bases and Titrations

- Balance the following equations and write the corresponding complete ionic and net ionic equations (if appropriate).
  - $\text{HC}_2\text{H}_3\text{O}_2 (\text{aq}) + \text{KOH} (\text{aq}) \rightarrow$
  - A solution of carbonic acid reacts with a solution of sodium hydroxide...
  - $\text{HNO}_3 (\text{aq}) + \text{Ba}(\text{OH})_2 (\text{aq}) \rightarrow$
- What volume of each of the following acids will completely react with 50.00 mL of 0.200 M NaOH?
  - 0.100 M HCl
  - 0.200 M  $\text{H}_3\text{PO}_4$  (hint: polyprotic acid)
- Complete the following two problems, which focus on neutralization reactions.
  - What volume of 0.115 M  $\text{HClO}_4$  solution is needed to neutralize 50.00 mL of 0.0875 M NaOH?
  - What volume of 0.128 M HCl is needed to neutralize 2.87 grams of  $\text{Mg}(\text{OH})_2$ ?
- A student titrates an unknown amount of potassium hydrogen phthalate ( $\text{KHC}_8\text{H}_4\text{O}_4$ , often abbreviated as KHP) with 20.46 mL of a 0.1000 M NaOH solution. KHP (molar mass = 204.22 g/mol) has one acidic hydrogen. What mass of KHP was titrated (reacted completely) by the sodium hydroxide solution?
- Tartaric acid,  $\text{H}_2\text{C}_4\text{H}_4\text{O}_6$ , has two acidic hydrogens. The acid is often present in wines and precipitates from solution as the wine ages. A solution containing an unknown concentration of the acid is titrated with NaOH. It requires 24.65 mL of 0.2500 M NaOH solution to titrate both acidic protons in 50.00 mL of the tartaric acid solution. Write a balanced net ionic equation for the neutralization reaction equation, and calculate the molarity of the tartaric acid solution.
- A 0.5895 gram sample of impure magnesium hydroxide is dissolved in a 100.0 mL of 0.2050 M HCl. The excess acid then needs 19.85 mL of 0.1020 M NaOH for neutralization. Calculate the percent by mass of magnesium hydroxide in the sample, assuming that it is the only substance reacting with the HCl solution.

# Chemistry 1 Stoichiometry Problem Set 3 Answers

**Ralph A. Burns, John William Hill**



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## **Table of Contents Chemistry 1 Stoichiometry Problem Set 3 Answers**

1. Understanding the eBook Chemistry 1 Stoichiometry Problem Set 3 Answers
  - The Rise of Digital Reading Chemistry 1 Stoichiometry Problem Set 3 Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry 1 Stoichiometry Problem Set 3 Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry 1 Stoichiometry Problem Set 3 Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry 1 Stoichiometry Problem Set 3 Answers
  - Personalized Recommendations
  - Chemistry 1 Stoichiometry Problem Set 3 Answers User Reviews and Ratings
  - Chemistry 1 Stoichiometry Problem Set 3 Answers and Bestseller Lists

5. Accessing Chemistry 1 Stoichiometry Problem Set 3 Answers Free and Paid eBooks
  - Chemistry 1 Stoichiometry Problem Set 3 Answers Public Domain eBooks
  - Chemistry 1 Stoichiometry Problem Set 3 Answers eBook Subscription Services
  - Chemistry 1 Stoichiometry Problem Set 3 Answers Budget-Friendly Options
6. Navigating Chemistry 1 Stoichiometry Problem Set 3 Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Chemistry 1 Stoichiometry Problem Set 3 Answers Compatibility with Devices
  - Chemistry 1 Stoichiometry Problem Set 3 Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemistry 1 Stoichiometry Problem Set 3 Answers
  - Highlighting and Note-Taking Chemistry 1 Stoichiometry Problem Set 3 Answers
  - Interactive Elements Chemistry 1 Stoichiometry Problem Set 3 Answers
8. Staying Engaged with Chemistry 1 Stoichiometry Problem Set 3 Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemistry 1 Stoichiometry Problem Set 3 Answers
9. Balancing eBooks and Physical Books Chemistry 1 Stoichiometry Problem Set 3 Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemistry 1 Stoichiometry Problem Set 3 Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chemistry 1 Stoichiometry Problem Set 3 Answers
  - Setting Reading Goals Chemistry 1 Stoichiometry Problem Set 3 Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry 1 Stoichiometry Problem Set 3 Answers
  - Fact-Checking eBook Content of Chemistry 1 Stoichiometry Problem Set 3 Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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