

Chemistry 1st semester final exam review Unit 1 & 2

Unit 1, Chapter 1 & 3 Introduction to Chemistry

1. Define the following words:

- Accuracy: how close a measurement is to the TRUE value (must have true value to measure)
- Precision: how close measurements are to each other (must have more than one measurement)
- Scientific Theory: well-tested explanation of some aspect of the natural world, attempts to explain why something occurs/happens
- Scientific Law: a concise statement that summarizes the results of many observations and experiments, does NOT attempt to explain why.

2. What is the difference between a scientific law and a scientific theory? A theory attempts to explain why, a law does not.

3. Describe the following measurement in terms of accuracy and precision: A laboratory analyzed a standard known to contain 140 ppb lead. The following results were obtained.

Trial	ppb Pb
1	169
2	114
3	142
4	115

The measurements were neither accurate nor precise. Only one measurement (trial 3) was close to the actual value. And only two measurements (Trial 2 & 4) were close to each other.

4. Describe the following measurement in terms of accuracy and precision:

- Measured value: 5.2 mg known value = 5.0 mg Accurate, NOT precise
- Measured value: 5.2 mg, 4.9 mg, 5.1 mg known value = 5.0 mg Both
- Measured values: 6.61 mL, 6.99 mL, 7.25 mL NOT Accurate, yet precise
- Measured value: 2.134 g/cm³ Neither accurate, nor precise

5. How many significant figures are in the following numbers?

Significant figure rules summarized:

numbers are always significant,

zeros in between significant numbers are significant

zeros at the END of the number AND after the decimal DO count

zeros at the end of a number WITH OUT a decimal don't count

zeros at the beginning of the number with a decimal don't count

1) 7100	2	2) 260.0	4	3) 0.00010	2	4) 218	3
5) 320	2	6) 0.00530	3	7) 22,568	5	8) 4,755.50	6

6. Express the Following in Scientific Notation:

- 0.000 033 3.3×10^{-5}
- 8 200 000 8.2×10^6
- 55 000 000 5.5×10^7
- 0.00288 2.88×10^{-3}
- 0.000 00733 7.33×10^{-6}
- 65 000 6.5×10^4

7. Convert the following metric measurements. Show your work:

- 1.6 cs to Gs $1.6 \text{ cs} \times \frac{10^{-8} \text{ s}}{1 \text{ cs}} \times \frac{1 \text{ Gs}}{10^9 \text{ s}} = 1.6 \times 10^{-11} \text{ Gs}$
- $8.8 \times 10^{-2} \text{ mm}$ to Mm $8.8 \times 10^{-2} \text{ mm} \times \frac{10^{-3} \text{ m}}{1 \text{ mm}} \times \frac{1 \text{ Mm}}{10^6 \text{ m}} = 8.8 \times 10^{-11} \text{ Mm}$
- $6.1 \times 10^{21} \text{ daJ}$ to TJ $6.1 \times 10^{21} \text{ daJ} \times \frac{10^{-1} \text{ J}}{1 \text{ daJ}} \times \frac{1 \text{ TJ}}{10^{12} \text{ J}} = 6.1 \times 10^{-2} \text{ TJ}$
- $9.5 \times 10^{-23} \text{ ML}$ to μL $9.5 \times 10^{-23} \text{ ML} \times \frac{10^6 \text{ L}}{1 \text{ ML}} \times \frac{1 \mu\text{L}}{10^{-6} \text{ L}} = 9.5 \times 10^{-11} \mu\text{L}$
- $2.35 \times 10^8 \text{ ng}$ to dag $2.35 \times 10^8 \text{ ng} \times \frac{10^{-9} \text{ g}}{1 \text{ ng}} \times \frac{1 \text{ dag}}{10^5 \text{ g}} = 2.35 \times 10^{-2} \text{ dag}$
- $1.78 \times 10^8 \text{ pg}$ to g $1.78 \times 10^8 \text{ pg} \times \frac{10^{-12} \text{ g}}{1 \text{ pg}} = 1.78 \times 10^{-4} \text{ g}$
- 274300 TJ to J $274300 \text{ TJ} \times \frac{10^{12} \text{ J}}{1 \text{ TJ}} = 2.743 \times 10^{17} \text{ J}$
- 0.00432 daJ to GJ $0.00432 \text{ daJ} \times \frac{10^1 \text{ J}}{1 \text{ daJ}} \times \frac{1 \text{ GJ}}{10^9 \text{ J}} = 4.32 \times 10^{-11} \text{ GJ}$

Chemistry 1st Semester Final Exam Review Unit 1 2

Oscar Krisen Buros



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