

Chemistry 2nd Semester Final Review **KEY**

This is meant as a guide. It does NOT contain every single thing you need to know for the final exam. Review notes, benchmark review sheets, homeworks, chapter study guides, etc.

Bring your book to school by Monday for turn in (seniors should bring book to bookstore by themselves ASAP)

Gas Laws (Ch. 13):

Subtopics: atmospheric pressure, Dalton's law of partial pressures, combined gas law, Ideal gas law, gas stoichiometry (22.4 L/mol for a gas at STP)

Practice questions:

1. If a gas occupies 3.8 L at a pressure of 2.71 atm, what would the pressure be if the volume changes to 1.47 L? (Temp. and amt. of gas constant)
 $V_1 = 3.8 \text{ L}; P_1 = 2.71 \text{ atm}$
 $V_2 = 1.47 \text{ L}; P_2 = ? \text{ Atm}$
Since all units are standard (L, atm), we do not need to convert the units before proceeding.
 $(P_1 V_1)/T_1 = (P_2 V_2)/T_2$
T goes away (not mentioned in problem)
 $P_1 V_1 = P_2 V_2$
Isolate for P_2 : $P_2 = (P_1 V_1)/V_2$
 $P_2 = (2.71 \text{ atm} \times 3.8 \text{ L})/(1.47 \text{ L}) = 7.0 \text{ atm}$ (L cancelled and it's 2 sig figs)
2. A gas has a volume of 1.49 L at a temperature of 34.75 degrees Celsius. What would the volume be at 78.41 degrees Celsius? (pressure & amt. of gas constant)
 $V_1 = 1.49 \text{ L}, T_1 = 34.75 \text{ degrees Celsius}$
 $V_2 = ? \text{ L}, T_2 = 78.41 \text{ degrees Celsius}$
We need to convert from Celsius to Kelvins before proceeding.
 $T_1 = 34.75 + 273 = 307.75 \text{ K}$
 $T_2 = 78.41 + 273 = 351.41 \text{ K}$
 $(P_1 V_1)/T_1 = (P_2 V_2)/T_2$
Get rid of P
 $V_1/T_1 = V_2/T_2$
Isolate V_2 : $V_2 = (V_1/T_1)T_2$
 $V_2 = (1.49 \text{ L}/307.75 \text{ K})(351.41 \text{ K}) = 1.70 \text{ L}$
3. What volume is occupied by 8.47 g of hydrogen gas at 84.7 degrees Celsius and 1.04 atm?
The combined gas law won't work here. You only have one condition and nothing changes.
 $PV = nRT$ will work. (n = moles and R = 0.08206 L atm/K mol)
 $P = 1.04 \text{ atm}, V = ? \text{ L}, n = \text{OOPS we need to convert from grams to moles!}, T = 84.7 \text{ degrees Celsius} + 273 = 357.7 \text{ K}$
 n : Hydrogen gas is diatomic!
 $8.47 \text{ g H}_2 \times (1 \text{ mol H}_2 / 2.02 \text{ g H}_2) = 4.193069307 \text{ mol H}_2$
Isolate V : $V = (nRT)/P = (4.193069307 \text{ mol H}_2 \times 0.08206 \text{ L atm/K mol} \times 357.7 \text{ K}) / 1.04 \text{ atm} = 118 \text{ L}$
4. What volume is occupied by 56.75 g of oxygen gas at STP?

Chemistry 2nd Semester Final Review Guide Answers

B Lingard



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Instructional Tests in Chemistry Earl Rouse Glenn, Louis E. Welton, 1929 *Nursing School Entrance Exam Review*
William H. Brown, Frederick A. Bettelheim, Jerry March, 2004-04 **Biology** Edward C. Gruber, 1963 *Playground and Recreation Director's Handbook* Arco Publishing Company, 1964 **Paperbound Books in Print**, 1992 **The Science Teacher**, 1968 SCC Library has 1964 cur The World Review, 1929 **Annual Report of the Illinois State Board of Health** Illinois State Board of Health, 1884 **Annual Report of Illinois State Board of Health** Illinois State Board of Health, 1883 Annual report of the State Board of Health of Illinois. 1884, 1884 **Conspectus of the Medical Colleges of America** Illinois State Board of Health, 1884 *Annual Catalogue* College of Hawaii, 1910 **The Software Encyclopedia 2001**, 2001 **The Journal of Educational Research**, 1921 *Interchange Level 1 Teacher's Edition with Assessment Audio CD/CD-ROM* Jack C. Richards, Jonathan Hull, Susan Proctor, 2012-08-09 Interchange Fourth Edition is a fully revised edition of Interchange the world's most successful series for adult and young adult learners of North American English The course has been revised to reflect the most recent approaches to language teaching and learning It remains the innovative series teachers and students have grown to love while incorporating suggestions from teachers and students all over the world This edition offers updated content in every unit grammar practice and opportunities to develop speaking and listening skills Interchange Fourth Edition features contemporary topics and a strong focus on both accuracy and fluency Its successful multi skills syllabus integrates themes grammar functions vocabulary and pronunciation The underlying philosophy of the course remains that language is best learned when it's used for meaningful communication **Atoms, Molecules and Reactions** Joseph D. Laposa, 1994 **The Detroit Educational Bulletin** Detroit (Mich.). Board of Education, 1926 Vols 2-7 contain also Special bulletins pub during the same period *The 2002 Guide to the Evaluation of Educational Experiences in the Armed Services* American Council on Education, 2002 Long considered to be the standard reference work in this area this three volume set describes more than 8 000 courses offered between January 1990 and the present by various service branches and the Department of Defense Long considered to be the standard reference work in this area this three volume set describes more than 8 000 courses offered between January 1990 and the present by various service branches and the Department of Defense Updated every two years Rocky Mountain Druggist, 1925 **Educational Review**, 1926

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