

# Redox reaction

No. of equivalents, equivalent mass, normality

Equivalent concept  $\rightarrow$  reaction independent

$\rightarrow$  reaction doesn't need to be balanced

Reaction  $\begin{cases} \rightarrow \text{Balanced} \rightarrow \text{use mole concept} \\ \rightarrow \text{Unbalanced} \begin{cases} \rightarrow \text{Balance it} \rightarrow \text{mole concept} \\ \rightarrow \text{Calculate } n_f \rightarrow \text{eq. concept} \end{cases} \end{cases}$

No. of equivalent  $\rightarrow$  no. of moles  $\times$  n-factor

$$\Sigma = \frac{\text{Molar mass}}{n_f}$$

$$\text{Normality} = \text{Molarity} \times n_f$$

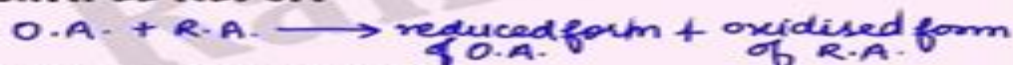
$$\Sigma = \frac{\text{Molar mass}}{\text{Total no. of } [H^+/\text{OH}^-] \text{ replaceable or to be replaced}}$$

is calculated for 1 mole only

$$N = \text{no. of equivalent/v}$$

Calculation of n-factor for different type of redox reaction-

$\rightarrow$  SIMPLE REDOX



$\rightarrow$  COMBINED REDOX



$\rightarrow$  SELF DECOMPOSITION REACTION



$$n_f = n \text{ Reduction}$$

In a single molecule, one element will undergo  $\text{oxi}^n$  & other one is reduced

$\rightarrow$  DISPROPORTIONATION REACTION



$$n_f = \frac{n_o \times n_R}{n_o + n_R}$$

same element gets oxidised as well as reduced



$$n = \frac{\text{total charge}}{\text{total item}} = \frac{2(7-2) + 1(7-4)}{3}$$

only for  $\text{rea}^n$  in which same element gets reduced in 2 diff. stages

$\rightarrow$  if oxidised  $\rightarrow$  combined redox

Redox reaction

it is basically transfer of  $e^-$



Law of equivalence

equivalent of acid = equivalent of base



Balancing of chemical reaction

- $\rightarrow$  Calculate n-factor of oxidising agent & reducing agent.
- $\rightarrow$  Cross multiply with the simplest ratio of n-f
- $\rightarrow$  Balance atoms under-going oxidation & reduction.
- $\rightarrow$  To Balance O & H

Acidic medium

O  $\rightarrow$  comes from  $\text{H}_2\text{O}$   
H  $\rightarrow$  from  $\text{H}^+$

Basic medium

O  $\rightarrow$  comes from  $\text{OH}^-$   
H  $\rightarrow$  from  $\text{H}_2\text{O}$



# Chemistry 2redox Reactions Study Guide

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## **Table of Contents Chemistry 2redox Reactions Study Guide**

1. Understanding the eBook Chemistry 2redox Reactions Study Guide
  - The Rise of Digital Reading Chemistry 2redox Reactions Study Guide
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry 2redox Reactions Study Guide
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry 2redox Reactions Study Guide
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry 2redox Reactions Study Guide
  - Personalized Recommendations
  - Chemistry 2redox Reactions Study Guide User Reviews and Ratings
  - Chemistry 2redox Reactions Study Guide and Bestseller Lists



5. Accessing Chemistry 2redox Reactions Study Guide Free and Paid eBooks
  - Chemistry 2redox Reactions Study Guide Public Domain eBooks
  - Chemistry 2redox Reactions Study Guide eBook Subscription Services
  - Chemistry 2redox Reactions Study Guide Budget-Friendly Options
6. Navigating Chemistry 2redox Reactions Study Guide eBook Formats
  - ePub, PDF, MOBI, and More
  - Chemistry 2redox Reactions Study Guide Compatibility with Devices
  - Chemistry 2redox Reactions Study Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemistry 2redox Reactions Study Guide
  - Highlighting and Note-Taking Chemistry 2redox Reactions Study Guide
  - Interactive Elements Chemistry 2redox Reactions Study Guide
8. Staying Engaged with Chemistry 2redox Reactions Study Guide
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemistry 2redox Reactions Study Guide
9. Balancing eBooks and Physical Books Chemistry 2redox Reactions Study Guide
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemistry 2redox Reactions Study Guide
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chemistry 2redox Reactions Study Guide
  - Setting Reading Goals Chemistry 2redox Reactions Study Guide
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry 2redox Reactions Study Guide
  - Fact-Checking eBook Content of Chemistry 2redox Reactions Study Guide
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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