

## Section Review

## Objectives

1. Explain the formation of ionic compounds.
2. Describe the properties of ionic compounds.

## Vocabulary

- |                   |                   |
|-------------------|-------------------|
| a. ionic compound | b. cation         |
| c. anion          | d. ionic compound |
| e. ionic bond     |                   |

## Part A Completion

Write the word or words that best complete each statement. Write complete sentences that will demonstrate your understanding of each statement. Be specific.

1. Ions are charged atoms or molecules. An atom that has lost one or more electrons is called a(n) cation. An atom that has gained one or more electrons is called a(n) anion. Ions are formed by the loss or gain of electrons from a neutral atom.

2. A(n) ionic compound is a compound that is made up of ions. It is formed by the combination of a cation and an anion. The ions are held together by a(n) ionic bond. The formula for sodium chloride is NaCl. The formula for calcium fluoride is CaF<sub>2</sub>.

## Part B True-False

Write true or false for each statement. If false, write the correct statement.

1. True The formula for sodium chloride is NaCl. The formula for calcium fluoride is CaF<sub>2</sub>.

# Chemistry 7 2 Ions Section Review Answers

**D Siedentop**



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