

**CHAPTER 14 REVIEW***Acids and Bases***SECTION 1****SHORT ANSWER** Answer the following questions in the space provided.

- 1.**
- Name the following compounds as acids:

_____ sulfuric acid _____	a. $\text{H}_2\text{SO}_4$
_____ sulfurous acid _____	b. $\text{H}_2\text{SO}_3$
_____ hydrosulfuric acid _____	c. $\text{H}_2\text{S}$
_____ perchloric acid _____	d. $\text{HClO}_4$
_____ hydrocyanic acid _____	e. hydrogen cyanide

- 2.**
- \_\_\_\_\_
- $\text{H}_2\text{S}$
- \_\_\_\_\_ Which (if any) of the acids mentioned in item
- 1**
- are binary acids?

- 3.**
- Write formulas for the following acids:

_____ $\text{HNO}_2$ _____	a. nitrous acid
_____ $\text{HBr}$ _____	b. hydrobromic acid
_____ $\text{H}_3\text{PO}_4$ _____	c. phosphoric acid
_____ $\text{CH}_3\text{COOH}$ _____	d. acetic acid
_____ $\text{HClO}$ _____	e. hypochlorous acid

- 4.**
- Calcium selenate has the formula
- $\text{CaSeO}_4$
- .

_____ $\text{H}_2\text{SeO}_4$ _____	a. What is the formula for selenic acid?
_____ $\text{H}_2\text{SeO}_3$ _____	b. What is the formula for selenous acid?

- 5.**
- Use an activity series to identify two metals that will not generate hydrogen gas when treated with an acid.

Choose from Cu, Ag, Au, Pt, Pd, or Hg.  
\_\_\_\_\_

- 6.**
- Write balanced chemical equations for the following reactions of acids and bases:

- a. aluminum metal with dilute nitric acid



- b. calcium hydroxide solution with acetic acid



# Describing Acids And Bases Section Review Answers

**Gabrielle Koutoukidis, Kate Stainton**



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