

Chapter 8: Chemical Equations and Reactions

I. Describing Chemical Reactions

A. A **chemical reaction** is the process by which one or more substances are changed into one or more different substances. A **chemical equation** represents, with symbols and formulas, the identities and relative amounts of the reactants and products in a chemical reaction.

B. Indications of a Chemical Reaction

- There are 4 indicators that show a reaction has taken place.

1. Evolution of heat and light.
2. Production of a gas.
3. Formation of a precipitate.
4. Color change.

C. Characteristics of Chemical Equations

1. There are 3 requirements for any chemical equations.

- a. The equation must represent known facts. All products and reactants must be identified.
- b. The equation must contain the correct formulas for the reactants and the products. Remember diatomics (H_2 , etc), sulfur (S_8), and phosphorous (P_4).
- c. The law of conservation of mass must be satisfied. To satisfy this law, coefficients are added where necessary.

2. Word and Formula Equations.

Describing Chemical Reactions Chapter 8 Answer Key

Lei Shi



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Chemistry Frank Brescia, 1978 *Numerical Techniques for Global Atmospheric Models* Peter H. Lauritzen, Christiane Jablonowski, Mark A. Taylor, Ramachandran D. Nair, 2011-03-29 This book surveys recent developments in numerical techniques for global atmospheric models It is based upon a collection of lectures prepared by leading experts in the field The chapters reveal the multitude of steps that determine the global atmospheric model design They encompass the choice of the equation set computational grids on the sphere horizontal and vertical discretizations time integration methods filtering and diffusion mechanisms conservation properties tracer transport and considerations for designing models for massively parallel computers A reader interested in applied numerical methods but also the many facets of atmospheric modeling should find this book of particular relevance **Chemical Thermodynamics: Principles and Applications** J. Bevan Ott, Juliana Boerio-Goates, 2000-06-07 *Chemical Thermodynamics Principles and Applications* presents a thorough development of the principles of thermodynamics an old science to which the authors include the most modern applications along with those of importance in developing the science and those of historical interest The text is written in an informal but rigorous style including anecdotes about some of the great thermodynamicists with some of whom the authors have had a personal relationship and focuses on real systems in the discussion and figures in contrast to the generic examples that are often used in other textbooks The book provides a basic review of thermodynamic principles equations and applications of broad interest It covers the development of thermodynamics as one of the pre eminent examples of an exact science A discussion of the standard state that emphasizes its significance and usefulness is also included as well as a more rigorous and indepth treatment of thermodynamics and discussions of a wider variety of applications than are found in more broadly based physical chemistry undergraduate textbooks Combined with its companion book *Chemical Thermodynamics Advanced Applications* the practicing scientist will have a complete reference set detailing chemical thermodynamics Outlines the development of the principles of thermodynamics including the most modern applications along with those of importance in developing the science and those of historical interest Provides a basic review of thermodynamic principles equations and applications of broad interest Treats thermodynamics as one of the preeminent examples of an exact science Provides a more rigorous and indepth treatment of thermodynamics and discussion of a wider variety of applications than are found in more

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