

$$.227 \text{ g H}_2\text{O} \frac{1 \text{ mol H}_2\text{O}}{18.02 \text{ g H}_2\text{O}} = 0.0126 \text{ mol H}_2\text{O}$$

0.0126 moles H₂O

2. Calculate the moles of copper in your sample of copper chloride?

$$0.742 \text{ g Cu} \frac{1 \text{ mol Cu}}{63.546 \text{ g Cu}} = 0.0117 \text{ mol Cu} \quad \underline{0.0117} \text{ moles Cu}$$

3. Calculate the moles of chlorine in your sample of copper chloride?

$$M_{\text{Cl}} = M_{\text{hydrate}} - M_{\text{Cu}} = 0.777 \text{ g} - 0.742 \text{ g} = 0.035 \text{ g Cl} \quad \underline{0.0005} \text{ moles Cl}$$

$$0.035 \text{ g Cl} \frac{1 \text{ mol Cl}}{70.906 \text{ g Cl}} = 0.0005 \text{ mol Cl}$$

4. Write the proper chemical formula for the compound that you tested.

$$\text{Moles Cu} = \frac{.0117}{.0005} = 2.3 \div 3 = 7$$

$$\text{Moles Cl} = \frac{.0005}{.0005} = 1 \div 3 = 3$$

Correct formula for initial hydrated salt _____

5. Explain using chemical equations the process of a) dehydrating the hydrate salt b) dissolving the dehydrated salt c) the change in copper observed after adding the aluminum wire to the solution.

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