

Empirical Formula of Epsom Salt

Introduction: Sometimes salts combine chemically with water and form hydrates in which a definite number of water molecules combine with the ions of a salt to form a crystal. Although chemically bound, this water of hydration maintains its characteristic composition in much the same manner as radicals do. For this reason we do not write $\text{H}_{20}\text{Na}_2\text{SO}_{14}$, but we write $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$. The dot means that each SO_4^{2-} ion and every two Na^+ ions are associated with 10 molecules of water. Some salts have more than one hydrate; i.e., sodium sulfate forms $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ and $\text{Na}_2\text{SO}_4 \cdot 7\text{H}_2\text{O}$. Hydrates are usually not very stable and all the water may be removed from the hydrated crystal by heating, leaving the anhydrous salt behind.

Many hydrates lose their water of hydration without heating when they are exposed to dry air. These substances are said to be efflorescent. Some salts, which may or may not be hydrates, absorb water from the air until a concentrated solution of the salt is formed. Such salts are deliquescent. They are good drying agents and are very soluble in water. Some anhydrous salts such as CaCl_2 may absorb enough water from the air to first form the hydrated crystal and then absorb additional water to form a solution.

Many materials such as silk, wool, hair, and tobacco pick up water vapor from the air but do not absorb enough to form a solution. Such materials are hygroscopic.

In this experiment a hydrated salt will be heated to constant weight, removing the water of hydration. The loss of weight will be the weight of water present in the original salt. From the weight of the water in the sample; the molecular weight of water, the weight of the anhydrous salt, and the formula of the anhydrous salt (which your instructor will give to you), the percent of water in the hydrate, and the number of moles of water associated with one mole of the anhydrous salt will be calculated.

Procedure: Obtain a sample of Epsom salt from your instructor. Weigh a clean, dry evaporating dish to the nearest 0.01 gram. Empty the contents of the test tube into the evaporating dish and weigh again.

Place the dish and contents on a wire screen and heat with a low flame. (This is necessary to keep from losing any of the solid). Gradually increase the heat after it appears that most of the water has been removed. After heating the dish and contents for about 30 minutes, let it cool and weigh. Reheat the dish and contents for about 15 minutes with a full flame to drive off any water that remains. Cool and weigh again. If there was a loss of weight between the first and second weighing, the salt was not dry the first time and the process should be repeated until a constant weight is obtained.

Calculations:

1. Calculate the percent of water in the hydrate.

$$\% \text{H}_2\text{O} = (\text{mass of H}_2\text{O} / \text{mass of epsom salt}) * 100$$

2. The anhydrous salt is magnesium sulfate, MgSO_4 . Calculate the formula of the hydrated salt.

a) **Molar Mass of MgSO_4** = Atomic mass of Mg + Atomic mass of S + 4 (Atomic mass of O)

b. **moles of MgSO_4** = Mass of MgSO_4 / molar mass of MgSO_4

c. **moles of H_2O** = mass of water / molar mass of water

d. The number of water molecules per anhydrous salt crystal = larger # of moles / smaller # of moles

Determination Of Epsom Salt Formula

**Cresacre George Moor, William
Partridge**



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