

**CHAPTER 11, Chemical Reactions** *(continued)*

9. Use the symbols in Table 11.1 on page 323 to write a skeleton equation for the following chemical reaction. Hydrochloric acid reacts with zinc to produce aqueous zinc(II) chloride and hydrogen gas.

**► Balancing Chemical Equations** (pages 324–328)

10. What is the law of conservation of mass?

In any physical or chemical change, mass is neither created nor destroyed.

11. Complete the flowchart for balancing equations.

Determine the correct formulas and physical states for the \_\_\_\_\_ **reactants** \_\_\_\_\_ and \_\_\_\_\_ **products** \_\_\_\_\_.

Write a \_\_\_\_\_ **skeleton equation** \_\_\_\_\_ with the formulas for the reactants on the left and the formulas for the products on the right of a yields sign ( $\longrightarrow$ ).

Count the number of \_\_\_\_\_ **atoms** \_\_\_\_\_ of each element in the reactants and in the products.

Balance the number of atoms of the elements on the two sides of the equation by placing \_\_\_\_\_ **coefficients** \_\_\_\_\_ in front of formulas. Never try to balance an equation by changing the \_\_\_\_\_ **subscripts** \_\_\_\_\_ in formulas.

Check each atom or polyatomic ion to be sure the equation is \_\_\_\_\_ **balanced** \_\_\_\_\_, and make sure that all coefficients are in the \_\_\_\_\_ **lowest** \_\_\_\_\_ possible ratio.

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