

Pre-Lab for Determination of Solubility Product Constant

1. Write the solubility product constant (K_{sp}) for the following reaction as a function of the concentrations of the products:



2. $\text{Ca(IO}_3)_2$ will ionize in water to produce IO_3^- ions. The IO_3^- ions will react with KI. Write the reaction for this reaction.
3. We will be using starch as an indicator. Why?
4. In this experiment, we will first produce I_2 by mixing calcium iodate with excess amount of KI. Based on the following reactions, how many moles of I_2 will be produced if we started with 6 moles of calcium iodate and excess amount of KI?



5. In this experiment, we will titrate I_2 with $\text{Na}_2\text{S}_2\text{O}_3$ (sodium thiosulfate). The reaction is provided below. If we needed 22.5 mL of 0.15 M $\text{Na}_2\text{S}_2\text{O}_3$ to reach the endpoint, how many moles of I_2 were present in the solution? **(Must show all calculation steps.)**



6. How many moles of $\text{Ca(IO}_3)_2$ are needed to produce the amount of I_2 in Problem #5?
7. The solubility product constant (K_{sp}) for calcium iodate is 6.47×10^{-6} . The molar mass of calcium iodate is 389.88 g/mol. 175.4 g of calcium iodate is added to 1 liter of water. What would be the concentrations of Ca^{2+} and IO_3^- at equilibrium? Would this solution form a precipitate? Is the solution unsaturated, saturated, or supersaturated when it reaches equilibrium? **(Must show all calculation steps.)**

Determination Of A Solubility Product Lab Answers

J Spring



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