

## Determining Empirical Formulas

|                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                     |         |                          |
|-------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------|---------|--------------------------|
| 1) 75% carbon, 25% hydrogen                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                         | Answer: | $\text{CH}_4$            |
| Work:<br>$\text{C } 75\% \rightarrow 75\text{g C} \left  \frac{1 \text{ mol}}{12.01\text{g C}} = 6.24 \text{ mol} / 6.24 = 1 \right.$<br>$\text{H } 25\% \rightarrow 25\text{g H} \left  \frac{1 \text{ mol H}}{1.01\text{g H}} = 24.75 \text{ mol} / 6.24 = 4 \right.$ <div style="float: right; border: 1px solid black; padding: 5px;"><math>\text{CH}_4</math></div>                                                                                                                                                            |         |                          |
| 2) 52.7% potassium, 47.3% chlorine                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                  | Answer: | $\text{KCl}$             |
| Work:<br>$\text{K } 52.7\% \rightarrow 52.7\text{g K} \left  \frac{1 \text{ mol K}}{39.10\text{g K}} = 1.34 \text{ mol} / 1.33 = 1 \right.$<br>$\text{Cl } 47.3\% \rightarrow 47.3\text{g Cl} \left  \frac{1 \text{ mol Cl}}{35.45\text{g Cl}} = 1.33 \text{ mol} / 1.33 = 1 \right.$ <div style="float: right; border: 1px solid black; padding: 5px;"><math>\text{KCl}</math></div>                                                                                                                                               |         |                          |
| 3) 22.1% aluminum, 25.4% phosphorous, 52.5 % oxygen                                                                                                                                                                                                                                                                                                                                                                                                                                                                                 | Answer: | $\text{AlPO}_4$          |
| Work:<br>$\text{Al } 22.1\% \rightarrow 22.1\text{g Al} \left  \frac{1 \text{ mol Al}}{26.98\text{g Al}} = 0.819 \text{ mol} / 0.819 = 1 \right.$<br>$\text{P } 25.4\% \rightarrow 25.4\text{g P} \left  \frac{1 \text{ mol P}}{30.97\text{g P}} = 0.820 \text{ mol} / 0.819 = 1 \right.$<br>$\text{O } 52.5\% \rightarrow 52.5\text{g O} \left  \frac{1 \text{ mol O}}{16.00\text{g O}} = 3.28 \text{ mol} / 0.819 = 4 \right.$ <div style="float: right; border: 1px solid black; padding: 5px;"><math>\text{AlPO}_4</math></div> |         |                          |
| 4) 13% magnesium, 87% bromine                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                       | Answer: | $\text{MgBr}_2$          |
| Work:<br>$13\text{g Mg} \left  \frac{1 \text{ mol Mg}}{24.31\text{g Mg}} = 0.54 \text{ mol Mg} / 0.54 = 1 \right.$<br>$87\text{g Br} \left  \frac{1 \text{ mol Br}}{79.90\text{g Br}} = 1.09 \text{ mol Br} / 0.54 = 2 \right.$ <div style="float: right; border: 1px solid black; padding: 5px;"><math>\text{MgBr}_2</math></div>                                                                                                                                                                                                  |         |                          |
| 5) 32.4% sodium, 22.5% sulfur, 45.1% oxygen                                                                                                                                                                                                                                                                                                                                                                                                                                                                                         | Answer: | $\text{Na}_2\text{SO}_4$ |
| Work:<br>$32.4\text{g Na} \left  \frac{1 \text{ mol}}{22.99\text{g Na}} = 1.41 \text{ mol Na} / 0.70 = 2 \right.$<br>$22.5\text{g S} \left  \frac{1 \text{ mol S}}{32.37\text{g S}} = 0.70 \text{ mol S} / 0.70 = 1 \right.$<br>$45.1\text{g O} \left  \frac{1 \text{ mol O}}{16\text{g O}} = 2.82 \text{ mol O} / 0.70 = 4 \right.$ <div style="float: right; border: 1px solid black; padding: 5px;"><math>\text{Na}_2\text{SO}_4</math></div>                                                                                    |         |                          |

# Determining Chemical Formulas Answer Key

**Howard Stephen Stoker**



## **Determining Chemical Formulas Answer Key:**

## Enjoying the Melody of Term: An Mental Symphony within **Determining Chemical Formulas Answer Key**

In a world used by displays and the ceaseless chatter of immediate conversation, the melodic elegance and mental symphony created by the published term frequently disappear into the back ground, eclipsed by the constant noise and disruptions that permeate our lives. But, set within the pages of **Determining Chemical Formulas Answer Key** a marvelous fictional prize brimming with fresh feelings, lies an immersive symphony waiting to be embraced. Crafted by a masterful composer of language, that fascinating masterpiece conducts readers on a psychological journey, well unraveling the concealed songs and profound influence resonating within each cautiously constructed phrase. Within the depths of the moving examination, we will discover the book is key harmonies, analyze its enthralling publishing design, and surrender ourselves to the profound resonance that echoes in the depths of readers souls.

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### **Table of Contents Determining Chemical Formulas Answer Key**

1. Understanding the eBook Determining Chemical Formulas Answer Key
  - The Rise of Digital Reading Determining Chemical Formulas Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Determining Chemical Formulas Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Determining Chemical Formulas Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Determining Chemical Formulas Answer Key
  - Personalized Recommendations

- Determining Chemical Formulas Answer Key User Reviews and Ratings
- Determining Chemical Formulas Answer Key and Bestseller Lists
- 5. Accessing Determining Chemical Formulas Answer Key Free and Paid eBooks
  - Determining Chemical Formulas Answer Key Public Domain eBooks
  - Determining Chemical Formulas Answer Key eBook Subscription Services
  - Determining Chemical Formulas Answer Key Budget-Friendly Options
- 6. Navigating Determining Chemical Formulas Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Determining Chemical Formulas Answer Key Compatibility with Devices
  - Determining Chemical Formulas Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Determining Chemical Formulas Answer Key
  - Highlighting and Note-Taking Determining Chemical Formulas Answer Key
  - Interactive Elements Determining Chemical Formulas Answer Key
- 8. Staying Engaged with Determining Chemical Formulas Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Determining Chemical Formulas Answer Key
- 9. Balancing eBooks and Physical Books Determining Chemical Formulas Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Determining Chemical Formulas Answer Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Determining Chemical Formulas Answer Key
  - Setting Reading Goals Determining Chemical Formulas Answer Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Determining Chemical Formulas Answer Key
  - Fact-Checking eBook Content of Determining Chemical Formulas Answer Key

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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