

## DETERMINING MOLECULAR FORMULAS (TRUE FORMULAS)

Solve the problems below

You may upload a picture of your work on paper, if you prefer...

1. The empirical formula of a compound is  $\text{NO}_2$ . Its molecular mass is 92 g/mol. What is its molecular formula?

2. The empirical formula of a compound is  $\text{CH}_2$ . Its molecular mass is 70 g/mol. What is its molecular formula?

3. A compound is found to be 40.0% carbon, 6.7% hydrogen, and 53.5% oxygen. Its molecular mass is 60 g/mol. What is its molecular formula?

4. A compound is 64.9% carbon, 13.5% hydrogen, and 21.6% oxygen. Its molecular mass is 74 g/mol. What is its molecular formula?

5. A compound is 54.5% carbon, 9.1% hydrogen, and 36.4% oxygen. Its molecular mass is 88 g/mol. What is its molecular formula?

# Determining Molecular Formulas True Formulas

**Charles Gilpin Cook**



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**A Course in General Chemistry** William McPherson, William Edwards Henderson, 1921

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## **Table of Contents Determining Molecular Formulas True Formulas**

1. Understanding the eBook Determining Molecular Formulas True Formulas
  - The Rise of Digital Reading Determining Molecular Formulas True Formulas
  - Advantages of eBooks Over Traditional Books
2. Identifying Determining Molecular Formulas True Formulas
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Determining Molecular Formulas True Formulas
  - User-Friendly Interface
4. Exploring eBook Recommendations from Determining Molecular Formulas True Formulas
  - Personalized Recommendations
  - Determining Molecular Formulas True Formulas User Reviews and Ratings
  - Determining Molecular Formulas True Formulas and Bestseller Lists
5. Accessing Determining Molecular Formulas True Formulas Free and Paid eBooks

- Determining Molecular Formulas True Formulas Public Domain eBooks
- Determining Molecular Formulas True Formulas eBook Subscription Services
- Determining Molecular Formulas True Formulas Budget-Friendly Options
- 6. Navigating Determining Molecular Formulas True Formulas eBook Formats
  - ePub, PDF, MOBI, and More
  - Determining Molecular Formulas True Formulas Compatibility with Devices
  - Determining Molecular Formulas True Formulas Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Determining Molecular Formulas True Formulas
  - Highlighting and Note-Taking Determining Molecular Formulas True Formulas
  - Interactive Elements Determining Molecular Formulas True Formulas
- 8. Staying Engaged with Determining Molecular Formulas True Formulas
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Determining Molecular Formulas True Formulas
- 9. Balancing eBooks and Physical Books Determining Molecular Formulas True Formulas
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Determining Molecular Formulas True Formulas
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Determining Molecular Formulas True Formulas
  - Setting Reading Goals Determining Molecular Formulas True Formulas
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Determining Molecular Formulas True Formulas
  - Fact-Checking eBook Content of Determining Molecular Formulas True Formulas
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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

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