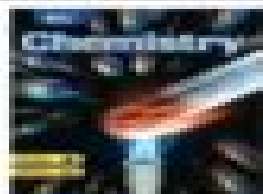


# Unit 6: Atomic and electronic structure

## reading



history of atomic models.....	77-88
atomic models, spectra.....	88-94
electron configuration.....	94-99
periodic table: trends.....	115-122
periodic trends.....	132-143

## major topics

1. An improved, modern atomic model
2. Configuration of electrons in thin atoms
3. Evidence for the relative energy of electrons in each atomic orbital
4. Factors and reasons for periodic variations in atomic properties

## objectives

6.1 Define, apply and distinguish among the following key terms:

- |                      |                           |                           |
|----------------------|---------------------------|---------------------------|
| 1. proton            | 8. electron configuration | 15. family of elements    |
| 2. electron          | 9. orbital configuration  | 16. group of elements     |
| 3. neutron           | 10. energy level          | 17. alkali metals         |
| 4. isotope           | 11. ground state          | 18. alkaline earth metals |
| 5. atomic number     | 12. excited state         | 19. transition elements   |
| 6. mass number       | 13. ionization energy     | 20. noble gases           |
| 7. isotopic notation | 14. valence electrons     | 21. halogens              |

6.2 Determine the number of protons, neutrons and electrons in any atom given its atomic number, mass number, and charge and vice versa.

6.3 Determine the electron configuration for any atom or ion in its ground state.

6.4 State the general trends in ionization energies, electronegativity, and atomic size across rows and down columns in the periodic table.

6.5 State the relationship between the size of a neutral atom and the size of that atom's cations or anions.

6.6 Explain trends in an atom's first ionization energy and its successive ionization energies based on its electron configuration and its position in the periodic table.

6.7 Explain trends in atoms' electronegativity and sizes based on their electron configurations and their positions in the periodic table.

6.8 Relate an atom's valence electron configuration to its typical ionic state and its position in the periodic table.

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**Shasha Hu**

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