

General chemistry 1 cumulative final exam

What is the sum of all coefficients after balancing the following equation:
 $\text{C}_{12}\text{H}_8\text{N}_2 + \text{O}_2 + \text{OH}^- \longrightarrow \text{CO}_3^{2-} + \text{H}_2\text{O} + \text{N}_2$ - Answer 68

What is the coefficient for O_2 after balancing the following equation?
 $\text{O}_2 + \text{C}_{10}\text{H}_{22} \longrightarrow \text{CO}_2 + \text{H}_2\text{O}$ - Answer 31

Acetylene torches burn acetylene by combining it with oxygen gas to produce carbon dioxide and water according to the following balanced equation:



Use the following molar masses:

O_2 is 32.00 g/mol, CO_2 is 44.01 g/mol, H_2O is 18.02 g/mol

If there are 208 g acetylene and 647 g of oxygen gas, which will be the limiting reactant? - Answer C_2H_2

If only 558 g of CO_2 were actually produced in the above reaction ($2 \text{C}_2\text{H}_2 + 5 \text{O}_2 \longrightarrow 4 \text{CO}_2 + 2\text{H}_2\text{O}$), what is the percent yield of the reaction? - Answer 79.4%

How many milliliters of 0.200 M NH_4OH are needed to react with 12.0 mL of 0.550 M FeCl_3 ?



Choose the pair where the formula of the acid is correctly matched with the name.

H_2S - hydrosulfuric acid

HNO_3 - nitrous acid

H_2CO_3 - hydrocarbonic acid

HClO_2 - chloric acid - Answer H_2S - hydrosulfuric acid

What is the name of the compound: $\text{Fe}_3(\text{PO}_4)_2$?

Iron(III) phosphate

Iron(II) phosphate

Iron phosphate

Triiron diphosphate - Answer Iron(II) phosphate

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JA Banks



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