



5 En effet, une minute après, Pritchard sortait de la gare, poursuivi par la servante armée d'un balai. Il tenait à la gueule une côtelette qu'il venait de prendre sur le gril.

6 « Que va dire mon maître ? pleurait la femme.
— Il va dire que Pritchard aime la viande.
— Oui, mais il ne sera pas content.
— Ne craignez rien, il ne grondera pas... Et dites-lui que je l'emmène déjeuner chez Vatrin.
— C'est égal, il lui arrivera malheur, à votre chien. Je ne vous dis que cela, il lui arrivera malheur ! »

EXERCICES.

1. Copie-écriture. — Copier le n° 5 de la lecture.

2. Copie-grammaire. — Copier les noms suivants puis les mettre au pluriel. (Ne pas oublier l's).

Ex. : Le domestique, les domestiques.

<i>le domestique</i>	<i>la plaine</i>	<i>le pont</i>	<i>le déjeuner</i>
<i>la descente</i>	<i>le chien</i>	<i>la dent</i>	<i>la côtelette</i>

3. Grammaire. — Copier les noms suivants. Suivant le cas mettre le ou la devant chacun d'eux. Ex. : La femme.

<i>femme</i>	<i>balai</i>	<i>main</i>	<i>minute</i>
<i>guêtre</i>	<i>soulier</i>	<i>chienne</i>	<i>servante</i>

Conte Dune Infideacuteliteacute Peu Ordinaire

Jin-Ying Zhang



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a net ionic equation is the most accurate representation of the actual chemical process that occurs writing these equations requires a familiarity with solubility rules acid base reactivity

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apr 25 2017 this is understood when you consider the two chemicals appear unchanged before and after the reaction therefore the net ionic equation is $2\text{Cl}^- + 2\text{Ag}^+$

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ions that remain essentially unchanged during a reaction are called spectator ions this means you can ignore them when you write the ionic equation you only need to model how the solid

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the first step to writing a net ionic equation is balancing the chemical equation present let s use the reaction between sodium chloride and silver nitrate as an example next we write the

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highlights learning objectives by the end of this section you will be able to derive chemical equations from narrative descriptions of chemical reactions write and balance chemical

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the net ionic equation is the chemical equation that shows only those elements compounds and ions that are directly involved in the chemical reaction notice that in writing the net ionic

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this net ionic equation indicates that solid silver chloride may be produced from dissolved chloride and silver I^+ ions regardless of the source of these ions these molecular and

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write the overall chemical equation the complete ionic equation and the net ionic equation for the reaction of aqueous silver fluoride with aqueous sodium phosphate to give solid silver

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the complete ionic equation for this reaction is as follows equation 4 13 $2\text{Ag}^+\text{aq} + 2\text{F}^-\text{aq} + 2\text{NH}_4^+\text{aq} + \text{Cr}_2\text{O}_7^{2-}\text{aq} \rightarrow \text{Ag}_2\text{Cr}_2\text{O}_7\text{s} + 2\text{NH}_4^+\text{aq} + 2\text{F}^-\text{aq}$ because two NH_4^+ aq

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