# **Bond Energy**

What makes a reaction endothermic or exothermic?

#### Why?

Chemical reactions can be a lot like playing with Legos\*—you must take apart part of your last creation before you can teplace it with something new. For many chemical reactions, we have to first break bonds in the reactants before we can put the atoms back into a new astrongement to form the products. Both of these processes involve changes in energy. The net energy change for a reaction is called the heat of reaction or the change in enthalpy (AH). In this activity we will look at one way energy changes can be approximated for chemical reactions.

# Model 1 - Breaking and Forming Bonds

	Reaction	Change in Energy (kJ/mod_)		Reaction	Change in Energy (kJ/mol)
A)	$\mathrm{PCl}_{\mathfrak{g}}(\mathfrak{g}) \ \rightarrow \ \mathbb{P}(\mathfrak{g}) \ + \ 3\mathrm{Cl}(\mathfrak{g})$	+ 966.7 broken	E)	$P(g) + 3Cl(g) \rightarrow PCl_{g}(g)$	- 966.7 formed
B)	$PCl_{g}(g) \rightarrow P(g) + 5Cl(g)$	+ 1.297.9 broken	F)	$P(g) + 5O(g) \rightarrow PO(g)$	- 1,297.9 formed
C)	$PF_3(g) \rightarrow P(g) + 3F(g)$	+ 1,470.4 broken	(2)	$P(g) + 3F(g) \rightarrow PF_3(g)$	- 1,470.4 formed
ED)	$PF_3(g) \rightarrow P(g) + 5F(g)$	+ 2,305.4 broken	H)	P(g) + 5F(g) → PF <sub>3</sub> (g)	- 2,305.4 formed

- Identify four reactions in Model 1 where bonds are being broken. Write "broken" in the Change in Energy box in the table for each reaction you identify.
- Identify four reactions in Model 1 where bonds are being formed. Write "formed," in the Change in Energy box in the table for each reaction you identify.
- 3. Circle the correct word to complete each sentence below,
  - at. When bonds are (broken) formed) there is a positive energy change.
  - 6. Breaking bunds is (endothermic).
  - c. When honds are (broken/formed) there is a negative energy change.
  - al. Forming bonds is (endothermic/exothermic).
- Find two reactions in Model 1 that are exact opposites of each other, that is, one reaction is the reverse of the other reaction.
  - a. How do the changes in energy for the reverse reactions compare?
    The changes in energy for these reactions are the same, but opposite in sign.
  - Explain your answer to part a considering what you learned from Questions 1—3 about bond breaking and bond formation.

If a given amount of energy is required to break a bond, then the same amount of energy will be released when the bond is formed.

# **Bond Energy Graph Pogil**

**Lingsheng Yao** 



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#### **Table of Contents Bond Energy Graph Pogil**

- 1. Understanding the eBook Bond Energy Graph Pogil
  - The Rise of Digital Reading Bond Energy Graph Pogil
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Bond Energy Graph Pogil
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Bond Energy Graph Pogil
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Bond Energy Graph Pogil

- Personalized Recommendations
- Bond Energy Graph Pogil User Reviews and Ratings
- Bond Energy Graph Pogil and Bestseller Lists
- 5. Accessing Bond Energy Graph Pogil Free and Paid eBooks
  - Bond Energy Graph Pogil Public Domain eBooks
  - Bond Energy Graph Pogil eBook Subscription Services
  - Bond Energy Graph Pogil Budget-Friendly Options
- 6. Navigating Bond Energy Graph Pogil eBook Formats
  - o ePub, PDF, MOBI, and More
  - Bond Energy Graph Pogil Compatibility with Devices
  - Bond Energy Graph Pogil Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Bond Energy Graph Pogil
  - Highlighting and Note-Taking Bond Energy Graph Pogil
  - o Interactive Elements Bond Energy Graph Pogil
- 8. Staying Engaged with Bond Energy Graph Pogil
  - o Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Bond Energy Graph Pogil
- 9. Balancing eBooks and Physical Books Bond Energy Graph Pogil
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Bond Energy Graph Pogil
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Bond Energy Graph Pogil
  - Setting Reading Goals Bond Energy Graph Pogil
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Bond Energy Graph Pogil

- Fact-Checking eBook Content of Bond Energy Graph Pogil
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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