

## Bond Energy

What makes a reaction endothermic or exothermic?

### Why?

Chemical reactions can be a lot like playing with Legos®—you must take apart part of your last creation before you can replace it with something new. For many chemical reactions, we have to first break bonds in the reactants before we can put the atoms back into a new arrangement to form the products. Both of these processes involve changes in energy. The net energy change for a reaction is called the **heat of reaction** or the **change in enthalpy** ( $\Delta H$ ). In this activity we will look at one way energy changes can be approximated for chemical reactions.

### Model 1 – Breaking and Forming Bonds

	Reaction	Change in Energy (kJ/mol <sub>rxn</sub> )		Reaction	Change in Energy (kJ/mol <sub>rxn</sub> )
A)	$\text{PCl}_3(\text{g}) \rightarrow \text{P}(\text{g}) + 3\text{Cl}(\text{g})$	+ 966.7	E)	$\text{P}(\text{g}) + 3\text{Cl}(\text{g}) \rightarrow \text{PCl}_3(\text{g})$	– 966.7
B)	$\text{PCl}_5(\text{g}) \rightarrow \text{P}(\text{g}) + 5\text{Cl}(\text{g})$	+ 1,297.9	F)	$\text{P}(\text{g}) + 5\text{Cl}(\text{g}) \rightarrow \text{PCl}_5(\text{g})$	– 1,297.9
C)	$\text{PF}_3(\text{g}) \rightarrow \text{P}(\text{g}) + 3\text{F}(\text{g})$	+ 1,470.4	G)	$\text{P}(\text{g}) + 3\text{F}(\text{g}) \rightarrow \text{PF}_3(\text{g})$	– 1,470.4
D)	$\text{PF}_5(\text{g}) \rightarrow \text{P}(\text{g}) + 5\text{F}(\text{g})$	+ 2,305.4	H)	$\text{P}(\text{g}) + 5\text{F}(\text{g}) \rightarrow \text{PF}_5(\text{g})$	– 2,305.4

- Identify four reactions in Model 1 where bonds are being broken. Write "broken" in the Change in Energy box in the table for each reaction you identify.
- Identify four reactions in Model 1 where bonds are being formed. Write "formed" in the Change in Energy box in the table for each reaction you identify.
- Circle the correct word to complete each sentence below.
  - When bonds are (broken/formed) there is a positive energy change.
  - Breaking bonds is (endothermic/exothermic).
  - When bonds are (broken/formed) there is a negative energy change.
  - Forming bonds is (endothermic/exothermic).
- Find two reactions in Model 1 that are exact opposites of each other, that is, one reaction is the reverse of the other reaction.
  - How do the changes in energy for the reverse reactions compare?
  - Explain your answer to part *a* considering what you learned from Questions 1–3 about bond breaking and bond formation.

# Bond Energy Pogil Model 4

**R Bogdan**

A red circular graphic with a gradient, appearing as a stylized arrow or a decorative element, located to the right of the author's name.

## **Bond Energy Pogil Model 4:**

## Unveiling the Power of Verbal Beauty: An Mental Sojourn through **Bond Energy Pogil Model 4**

In some sort of inundated with monitors and the cacophony of immediate connection, the profound energy and mental resonance of verbal artistry frequently diminish into obscurity, eclipsed by the regular onslaught of noise and distractions. Yet, nestled within the musical pages of **Bond Energy Pogil Model 4**, a fascinating work of literary brilliance that impulses with organic emotions, lies an wonderful trip waiting to be embarked upon. Written by a virtuoso wordsmith, that enchanting opus guides readers on a mental odyssey, gently revealing the latent possible and profound affect embedded within the complex web of language. Within the heart-wrenching expanse of this evocative examination, we shall embark upon an introspective exploration of the book is key subjects, dissect its fascinating publishing type, and immerse ourselves in the indelible effect it leaves upon the depths of readers souls.

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