

**Study Guide: Chemical Bonding**

1. Atoms join together with other atoms in a process called \_\_\_\_\_ which forms new \_\_\_\_\_. This bonding takes place between \_\_\_\_\_ electrons in the \_\_\_\_\_ energy level.

2. Atoms want to have \_\_\_\_\_ electrons in their outer \_\_\_\_\_ level. These electrons are called \_\_\_\_\_. Group 1 elements have \_\_\_\_\_ valence electron while the Group 2 elements have \_\_\_\_\_. With Group 3-10, also known as the \_\_\_\_\_ metals, there is no relationship between the \_\_\_\_\_ and the number of valence electrons. With groups 11-12, \_\_\_\_\_ from the group number will give you the number of valence electrons for those \_\_\_\_\_.

3. There are three types of chemical \_\_\_\_\_. When ions with opposite charges bond together, it is called \_\_\_\_\_ bonding. A three dimensional shape that forms during this bonding is called a \_\_\_\_\_ with \_\_\_\_\_ bonding, electrons are shared between atoms. A \_\_\_\_\_ bond is formed between metal ions. A group of two or more atoms held together through chemical bonding is called a \_\_\_\_\_.

4. Write the number of valence electrons next to each element:

\_\_\_\_\_ Oxygen

\_\_\_\_\_ Chlorine

\_\_\_\_\_ Barium

\_\_\_\_\_ Sodium

\_\_\_\_\_ Xenon

\_\_\_\_\_ Plumbum

- outer
- chemical bonding
- valence
- substances

- transition
- eight
- two
- group number
- valence electrons
- one
- energy
- elements
- subtracting

- molecule
- covalent
- bonding
- metallic
- ionic
- crystal lattice

# Bonding Chemical Reactions Cloze Ing In On Science

**SA Adler**



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Chemical Bonds examines the nature of the chemical bonds answering fundamental questions about how they form how they  
are broken and how they help define life as we know it *Periodic Table and Chemical Bonding* Open University. S101

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compounds with new properties **Discovering Chemistry With Natural Bond Orbitals** Frank Weinhold,2012-06-15

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