

Stoichiometry Cheat Sheet

1 mole = molar mass of the substance	1 mole = 6.02×10^{23} particles	1 mole = 22.4L of a gas at STP
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$$\begin{array}{|c|} \hline \text{Amount given} \\ \hline \end{array}
 \times \frac{1 \text{ mol given}}{\begin{array}{|c|} \hline \text{a) Molar mass} \\ \text{b) } 6.02 \times 10^{23} \text{ particles} \\ \text{c) 22.4L} \\ \hline \end{array}}
 \times \frac{(y) \text{ mol sought}}{(x) \text{ mol given}}
 \times \frac{\begin{array}{|c|} \hline \text{d) Molar mass} \\ \text{e) } 6.02 \times 10^{23} \text{ particles} \\ \text{f) 22.4L} \\ \hline \end{array}}{1 \text{ mol sought}} = \text{ANSWER}$$

- a) – used when amount given is a mass (g)
 b) – used when amount given is particles (molecules, formula units, atoms)
 c) – used when amount given is liters of gas at STP (L)

 d) – used when answer is sought in mass (g)
 e) – used when answer is sought in particles (molecules, formula units, atoms)
 f) – used when answer is sought in liters of gas at STP (L)

(x) = coefficient of given substance from balanced equation
 (y) = coefficient of sought substance from balanced equation

Example: (mass to mass problem)

Given the equation $\text{Al}(\text{NO}_3)_3 + \text{NaOH} \rightarrow \text{Al}(\text{OH})_3 + \text{NaNO}_3$, find the mass of NaNO_3 formed if 275g of NaOH reacts.

- Given substance = NaOH
- Sought substance = NaNO_3
- Balance equation = $\text{Al}(\text{NO}_3)_3 + 3\text{NaOH} \rightarrow \text{Al}(\text{OH})_3 + 3\text{NaNO}_3$
- Start with given amount (275g NaOH) and solve for grams of NaNO_3 .

"3" mol of NaNO_3 , because coefficient of NaNO_3 in balanced equation is "3"

Molar mass of NaNO_3 , because question asks for "mass"

$$275\text{g NaOH} \times \frac{1 \text{ mol NaOH}}{39.997\text{g NaOH}} \times \frac{3 \text{ mol NaNO}_3}{3 \text{ mol NaOH}} \times \frac{84.994\text{g NaNO}_3}{1 \text{ mol NaNO}_3} = 584\text{g NaNO}_3$$

Molar mass of NaOH because question gives info "275g of NaOH "

"3" mol of NaOH because coefficient of NaOH in balanced equation is "3"

Clozeing In On Sciece Stoichiometry

SJ Ball



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Resources in education ,1988-07 *Resources in Education* ,1984 *The ETS Test Collection Catalog* Educational Testing Service. Test Collection,1993 The major source of information on the availability of standardized tests Wilson Library BulletinCovers commercially available standardized tests and hard to locate research instruments [Current Index to Journals in Education](#) ,1986 *Stoichiometry* Sydney Young,1918 This introduction was originally prefixed to Dr Findlay s Phase rule which was the first volume of the series issued It belongs properly however to this volume and is therefore included here Note p vii Introduction issued also separately in 1904 *Formulation and Stoichiometry* Emil J. Margolis,2012-12-06 The purpose of this book is to interpret more sensitively some of the offerings of the standard text book of general chemistry As a supplement thereto it covers various aspects of formulation and stoichiometry that are frequently treated far too perfunctorily or in many instances are not considered at all The inadequate attention often accorded by the comprehensive text to many topics within its proper purview arises understandably enough from the numerous broad and highly varied objectives set for the first year of the curriculum for modern chemistry in colleges and universities For the serious student this means more often than not the frustrations of questions unanswered The amplification that this book proffers in the immediate area of its subject covers the equations representing internal redox reactions not only of the simple but also of the multiple disproportionations of which the complexities often discourage an undertaking despite the challenge they offer distinctions to be observed in the balancing of equations in con trasting alkali basic and ammonia basic reaction media quantitative contributions made by the ionization or dissociation effects of electrolytes to the colligative properties of their solutions intensive application of the universal reaction principle of chemical equivalence to the stoichiometry of oxidation and reduction *Encyclopedia of Stoichiometry: Volume II (Biomedical Science)* Bruce Hopkin,2015-01-06 Stoichiometry deals with the statistical analysis of quantity of reacting agents and finished products in a chemical reaction The purpose of this text is to provide an insight on the significance of stoichiometry in the biomedical domain It brings a compilation of chosen research works and evaluations which offer latest updates related to stoichiometry at different stages In this book a range of topics like host guest chemistry with special emphasis on cyclodextrins and crown ethers derivatives matters relating to stoichiometry of metal complexes and lipids importance of stoichiometry in the identification of protein interactions experimental methods used in certain systems and the interrelation between biomedicine and environment leading to the formulation of the concept of biological stoichiometry have been presented This text will be of great significance providing essential scientific data to experts involved in associated fields **You Can Do Chemistry** Sunil Tanna,2018-11-19 A comprehensive guide to performing mole and stoichiometric calculations with numerous examples as well as questions and answers Covers calculations relating to solids solutions gases and electrolysis plus as limiting and excess reactants chemical yields atom economy and much more Fully up to date with the last international standards

including the revised definition of mole which was agreed on November 16th 2018 **A Stoichiometry Unit** David Callaghan,2004 Problems and Worked Examples in Chemistry Albert Holderness,John Lambert,1978 **Stoichiometry and Process Calculations** Mr. Rohit Manglik,2024-01-22 EduGorilla Publication is a trusted name in the education sector committed to empowering learners with high quality study materials and resources Specializing in competitive exams and academic support EduGorilla provides comprehensive and well structured content tailored to meet the needs of students across various streams and levels *Calculating with Stoichiometry* ,2009-01-01 Inquiries in Science Chemistry Series Calculating with Stoichiometry Teacher s Guide **Chemistry Skill Builder** Steven D. Gammon,1997 Chemistry SKILLBuilder gives students extra practice and feedback on three key topics nomenclature stoichiometry and balancing equations *Stoichiometry Unit Project* Luann Marie Decker,1998 Stoichiometry B. I. Bhatt,1996 **Mole Concepts and Stoichiometry** Dharshi Bopegedera,2017 *Chemistry Success in 20 Minutes a Day* Michael B. McGinnis,2005 Offers a diagnostic test and twenty lessons covering vital chemistry skills **Chemical Transformations in Inhibited Flames Over Range of Stoichiometry** O.P. Korobeinichev,A.G. Shmakov,V.M. Shvartsberg,2012 Chemistry for Students and Parents Roy Richard Sawyer,2013-12-15 If you are a parent struggling to help your child with chemistry homework this is a short book that will help you It covers key chemistry topics Oxides Bases Acids Salts Equivalent proportions Acid Base reactions Weight and Volume problems Equilibrium Le Chatelier s Principle Freezing and Boiling points Balance Redox Reactions 30 examples with explanations Stoichiometry 30 problems with answers and solutions If you are student read this book and you will prove to yourself that you can understand chemistry *Holt Chemistry 2* Holt Rinehart & Winston,Holt, Rinehart and Winston Staff,2003-01-01

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