

Experiment 21

Advance Study Assignment: Rates of Chemical Reactions, II. A Clock Reaction

1. A student studied the clock reaction described in this experiment. She set up Reaction Mixture 2 by mixing 20 mL 0.010 M KI, 10 mL 0.001 M $\text{Na}_2\text{S}_2\text{O}_8$, 10 mL 0.040 M KBrO_3 , and 10 mL 0.10 M HCl using the procedure given. It took about 45 seconds for the color to turn blue.

- a. She found the concentrations of each reactant in the reacting mixture by realizing that the number of moles of each reactant did not change when that reactant was mixed with the others, but that its concentration did. For any reactant A,

$$\text{no. moles A} = M_{A \text{ stock}} \times V_{\text{stock}} = M_{A \text{ mixture}} \times V_{\text{mixture}}$$

The volume of the mixture was 50 mL. Revising the above equation, she obtained

$$M_{A \text{ mixture}} = M_{A \text{ stock}} \times \frac{V_{\text{stock}} (\text{mL})}{50 \text{ mL}}$$

Find the concentrations of each reactant by using the above equation.

$$(\text{I}^-) = \text{_____ M}; (\text{BrO}_3^-) = \text{_____ M}; (\text{H}^+) = \text{_____ M}$$

- b. What was the relative rate of the reaction ($1000/t$)? _____
- c. Knowing the relative rate of reaction for Mixture 2 and the concentrations of I^- , BrO_3^- , and H^+ in that mixture, she was able to set up Equation 5 for the relative rate of the reaction. The only quantities that remained unknown were k' , m , n , and p . Set up Equation 5 as she did, presuming she did it properly. Equation 5 is on page 172.

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