

Exam 2013 Question 1 (a) Test for ions: Cl^- and Ag^+ ions

On your answer, for each ion, write the expected test, the expected observation, and the expected equation. For each ion, write the expected test, the expected observation, and the expected equation. For each ion, write the expected test, the expected observation, and the expected equation.

1. To test for Cl^- , add AgNO_3 to the solution. The solution turns white precipitate.
 Cl^- is present. $\text{AgNO}_3 + \text{Cl}^- \rightarrow \text{AgCl} \downarrow$
2. To test for Ag^+ , add HCl to the solution. The solution turns white precipitate.
 Ag^+ is present. $\text{Ag}^+ + \text{Cl}^- \rightarrow \text{AgCl} \downarrow$

Cl^- is present

3. The observed test is that the Cl^- is present. The presence of Cl^- is indicated by the solution turning white precipitate.
 $\text{AgNO}_3 + \text{Cl}^- \rightarrow \text{AgCl} \downarrow$

4. A positive test for Cl^- is present when the solution turns white precipitate. The observed test is that the Cl^- is present.
 $\text{AgNO}_3 + \text{Cl}^- \rightarrow \text{AgCl} \downarrow$

5. A positive test for Cl^- is present when the solution turns white precipitate. The observed test is that the Cl^- is present.
 $\text{AgNO}_3 + \text{Cl}^- \rightarrow \text{AgCl} \downarrow$

6. To test for Ag^+ , add HCl to the solution. The solution turns white precipitate.
 $\text{Ag}^+ + \text{Cl}^- \rightarrow \text{AgCl} \downarrow$

7. To test for Ag^+ , add HCl to the solution. The solution turns white precipitate.
 $\text{Ag}^+ + \text{Cl}^- \rightarrow \text{AgCl} \downarrow$

8. The test for Ag^+ is that the solution turns white precipitate. The observed test is that the Ag^+ is present.
 $\text{Ag}^+ + \text{Cl}^- \rightarrow \text{AgCl} \downarrow$

Chemical species to be tested	Test	Observation
Cl^-	Sample solution + AgNO_3	$\text{AgCl} \downarrow$ (white precipitate)
Ag^+	Sample solution + HCl	$\text{AgCl} \downarrow$ (white precipitate)
Cl^-	Sample solution + AgNO_3	$\text{AgCl} \downarrow$ (white precipitate)
Ag^+	Sample solution + HCl	$\text{AgCl} \downarrow$ (white precipitate)
Cl^-	Sample solution + AgNO_3	$\text{AgCl} \downarrow$ (white precipitate)
Ag^+	Sample solution + HCl	$\text{AgCl} \downarrow$ (white precipitate)

The ions present in the sample is AgCl

Chemical species to be tested	Test	Observation
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AgCl

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AgCl

The ions present in the sample is AgCl

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