

1. Determine the electron geometry (eg) and molecular geometry (mg) of SF_6 .
- A. eg=trigonal bipyramidal, mg=trigonal bipyramidal
 - B. eg=tetrahedral, mg=tetrahedral
 - C. **eg=trigonal bipyramidal, mg=see-saw**
 - D. eg=octahedral, mg=trigonal bipyramidal
 - E. eg=octahedral, mg=octahedral
2. How would you describe the molecular polarity of SeCl_6 ?
- A. Nonpolar, because Se-Cl is a nonpolar bond and there is no lone pair on Se
 - B. **Nonpolar, because Se-Cl is a polar bond but the geometry is symmetric**
 - C. Polar, because Se-Cl is a polar bond although the geometry is symmetric
 - D. Polar, because Se-Cl is a polar bond and there is a lone pair on Se
 - E. Polar, because Se-Cl is a nonpolar bond but there is a lone pair on Se
3. What is the hybridization of the central atom in BrF_5 ?
- A. sp
 - B. sp^2
 - C. sp^3
 - D. sp^3d
 - E. **sp^3d^2**
4. Describe a π bond.
- A. s orbital overlapping with d orbital
 - B. end to end overlap of two s orbitals
 - C. p orbital overlapping with f orbital
 - D. overlap of any two hybrid orbitals
 - E. **side by side overlap of two p orbitals**
5. Consider the Lewis structure for formic acid (CO_2H_2). Determine the hybridization of the carbon atom. Then determine the **total number of sigma (σ) and pi (π) bonds in the molecule.**
- A. Carbon hybridization= sp^3 , 3 σ bonds; 2 π bonds
 - B. Carbon hybridization= sp^3 , 4 σ bonds; 1 π bond
 - C. **Carbon hybridization= sp^2 , 4 σ bonds; 1 π bond**
 - D. Carbon hybridization= sp^2 , 3 σ bonds; 1 π bond
 - E. Carbon hybridization= sp^2 , 3 σ bonds; 2 π bonds

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