

Chemical and Physical Changes and Properties

Purpose

Based on experimental data, determine the occurrence of a chemical change.

Introduction

Text Reference: Chapter 1, *Chemistry, A Molecular Approach*, Nivaldo Tro, 4th ed.

Matter can experience two types of changes:

- A physical change is a change in physical properties (size, mass, state of matter, etc) without a change in chemical composition.
- A chemical change involves the rearrangement of atoms of one or more substances and a change in chemical properties or composition, resulting in the formation of at least one new substance. All chemical changes are accompanied by one or more physical change, along with energy exchanges.

A common example of a physical change is a change in state of the substance. The appearance of a substance changes when you melt, freeze, boil, condense or sublime it. This change is due to the addition or removal of heat. The composition of the substance remains the same, even if it looks very different. When heat is added to ice, it starts to melt and then the appearance is completely different, from a clear solid to a transparent liquid. If we continue adding heat, the liquid starts to vaporize, and the water now is steam. Even when the ice does not look very similar to steam, the composition is the same: it is water. Cutting, tearing, shattering, and grinding are other examples of physical changes. These may be irreversible, however there is no change in composition.

In a chemical change, a change in the composition occurs. For example, when the Statue of Liberty was erected in 1886, the copper metal was orange in color. However, by 1900, it had turned a turquoise color. The copper underwent a chemical change, going from metallic copper to a mixture of copper oxides, copper carbonates, copper sulfates, and copper chlorides due to exposure to various chemicals in the air and rain.

Certain observations that indicate a chemical change include:

- *Formation of precipitate.* It is the formation of an insoluble solid when two reactants are mixed. It may be a fine suspension or it may form large particles that quickly settle down. It may occur immediately, or it may take some time to settle down in order to be observed.
- *Color change.* It is the observation of a new color in the substance(s) that we are using. When a clean copper wire is heated in a burner flame, the appearance of the copper changes from coppery metallic to glowing red. Finally, its final appearance by cooling is a black solid. It is a new substance; copper oxide. However, when a clean platinum wire

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Mark A. Griep



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CHEM 111 Lab Manual General Chemistry Christofield, 2019-08-12 Laboratory Manual for Principles of General Chemistry J. A. Beran, Mark Lassiter, 2022-08-16 Laboratory Manual for Principles of General Chemistry 11th Edition covers two semesters of a general chemistry laboratory program The material focuses on the lab experiences that reinforce the concepts that not all experimental conclusions are the same and depend on identifying an appropriate experimental procedure selecting the proper apparatus employing the proper techniques systematically analyzing and interpreting the data and minimizing inherent variables As a result of good data a scientific and analytical conclusion is made which may or may not be right but is certainly consistent with the data Experiments write textbooks textbooks don't write experiments A student's scientific literacy grows when experiences and observations associated with the scientific method are encountered Further experimentation provides additional cause effect observations leading to an even better understanding of the experiment The 11th edition's experiments are informative and challenging while offering a solid foundation for technique safety and experimental procedure The reporting and analysis of the data and the pre and post lab questions focus on the intuitiveness of the experiment The experiments may accompany any general chemistry textbook and are compiled at the beginning of each curricular unit An Additional Notes column is included in each experiment's Report Sheet to provide a space for recording observations and data during the experiment Continued emphasis on handling data is supported by the Data Analysis section General Chemistry Lab Manual, CHM 125L Jean Smolen, 2011 *Acp Chm 1132 General Chemistry Lab Manual* Brooks/Cole, 2016-11-16 *Chem 111-112* Larry Alan Scheich, Tammy Janine Melton, New Mexico State University. Department of Chemistry and Biochemistry, 2002 Chem 106 Chemistry in Context 2 Mark A. Griep, 2017-01-03 **Custom CHEM 131/132** Brooks/Cole, 2015-07-20 **General College Chemistry I and II (CHM 11/12) Laboratory Manual** Darrell D. Ebbing, 2017 **Chemistry 132** Sharyl A. Majorski, 2012-08-15 General Chemistry Lab Manual Bobby Stanton, 2010 *Chemistry in Context Chem 106* Mark A. Griep, 2009-12-01 *Addison-Wesley Chemistry Laboratory Manual* Dennis D. Staley, 1995 *CHM141 General Chemistry I Lab Manual* Buyuktanir, 2015-06-09 **General Chemistry Lab Manual** Jan Gryko, Eugene Turner Knight, Jacksonville State University. Department of Physical and Earth Sciences, 2007 **General Chemistry Lab Manual 7E for Drexel University** Thorne, 2011-10-18 *Chemistry 1411* Jesse Yeh, 2012-07-20 **Chemistry Lab Manual** Spencer L. Seager, Michael R. Slabaugh, 2016 *Chemistry 23 Lab Manual - Outline of General Chem* Academx Publishing Services, Incorporated, 2012 *LAB MANUAL FOR CHEM 105* MARK A. GRIEP, 2022 Lab Manual for Principles of General Chemistry Jo Allan Beran, Benedictine University (Lisle, Ill.). Department of Chemistry/Biochemistry, 2011

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