

CHM 130

Lewis Dot Formulas and Molecular Shapes

Introduction

A chemical bond is an intramolecular (within the molecule) force holding two or more atoms together. Covalent chemical bonds are formed by valence electrons being shared between two different atoms. Both atoms attain the noble gas configuration with eight electrons (octet rule) or two electrons in their outer shell.

Octet Rule -

Atoms bond in such a way that each atom acquires eight electrons in its outer shell.

Duet Rule -

Hydrogen only requires 2 electrons to fill its outer shell and have He's electron configuration. ($1s^2$)

Lewis Dot Formula (also called an electron dot formula) -

Shows the valence electrons, indicating the bonding between atoms. The following guidelines will help draw the electron dot formulas correctly.

Guidelines for drawing simple Lewis dot formulas (electron dot formulas):

- (1) Draw a skeletal structure using single bonds to connect atoms to a central atom.
- (2) Calculate the total number of valence electrons.
- (3) Deduct one pair of valence electrons ($2e^-$ total) for each single bond drawn in Step #1, then use the remaining pairs to complete octets for the other atoms.
- (4) If there are not enough electrons to complete an octet for each atom, because the molecule is short $2e^-$, then move a nonbonding electron pair (an unshared pair) between two atoms that already share an electron pair to create a double bond. If the molecule is short $4e^-$, then move nonbonding pairs to create either two double bonds or one triple bond.
- (5) Check your work.

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Example 1: Draw the electron dot formula and build the molecular model of water, H_2O .

- (1) Draw a skeletal structure. Can hydrogen atoms be a central atom?
- (2) $H = 1$ valence electron $\times 2 = 2$ valence electrons; $O = 6$ valence electrons for a total of 8 valence electrons.
- (3) 8 valence electrons - 4 electrons committed in single bonds = 4 electrons to distribute. Place two electron pairs on O to complete its octet.
- (4) Check. There are 2 electrons on each H and there are 8 electrons around the central O atom.

Molecule	# of valence e^-	Lewis Structure
H_2O	$2(1) + 6 = 8e^-$	$H-\ddot{O}-H$

The bonds depicted in a Lewis dot formula (an electron dot formula) can form three types of bonds depending on the number of pairs of electrons shared between the two atoms forming the bond. The three types of bonds are:

Single Bond - sharing one pair of electrons ($2e^-$ total)

Double Bond - sharing two pairs of electrons ($4e^-$ total)

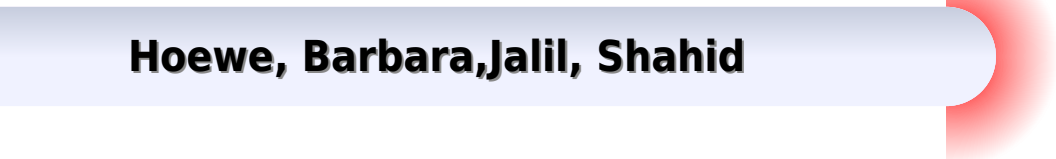
Triple bond - sharing three pairs of electrons ($6e^-$ total)

Example 2: Draw the electron dot formula and build the molecular model of sulfur trioxide, SO_3 .

- (1) Draw a skeletal structure with a central S atom.
- (2) Calculate the total number of valence electrons.
- (3) Deduct a pair of electrons for each of the single bonds drawn in Step #1. Use the remaining pairs to complete octets for each of the other atoms.
- (4) If there are not enough electrons to provide an octet for each atom, move a nonbonding (unshared) electron pair between two atoms that already share an electron pair.
- (5) Check your work. [Hint: Did you need a double bond?]

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Hoewe, Barbara, Jalil, Shahid



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