

CHM 001 Test Questions

Leveled Chapter Problems

1. What is the vapor density of R2SO4 (R=1, S=32, O=16) (a) 96 (b) 98 (c) 48 (d) 59
2. The atomic number is also called

(a) no. of protons (b) no. of electrons (c) no. of atoms (d) no. of neutrons

3. The Avogadro's number is (a) 6.023×10^{23} (b) 6.03×10^{23} (c) 6.022×10^{23} (d) 6.04×10^{23}
4. Find the number of moles of a substance present in 0.3g of a substance sample

(a) 0.27 moles (b) 0.26 moles (c) 0.24 moles (d) 0.25 moles

5. Molar mass is _____

(a) mass concentration/molar concentration (b) molar concentration/mass concentration (c) number of moles \times mass concentration (d) Molar ratio \times Molar Concentration

6. Which of these undergoes oxidation

(a) Ammonium Chloride (b) Oxygen Gas (c) Aluminium Oxide (d) Phosphorus Soap

7. Covalent bonds is found in (a) NaCl (b) Benzene Methyl (c) AlO (d) Oxygen Molecule

8. Ionic bond is found in (a) NH_4^+ (b) O=O (c) H-H (d) NaCl

9. Electronegativity based involves _____ while covalent involves _____ (a) Resonating, Sigma ring (b) Dangling, Triaxial (c) Triaxial, Sigma ring

10. Allotropes occur in the following Except _____ (a) C (b) S (c) Cl (d) P

11. _____ is NOT a diatomic (a) O (b) Cl (c) H (d) Al

12. The formula of Alkane is _____ (a) C_nH_{2n} (b) $\text{C}_n\text{H}_{2n-1}$ (c) $\text{C}_n\text{H}_{2n-2}$ (d) $\text{C}_n\text{H}_{2n+2}$

13. _____ is the most reactive element and why?

(a) He, Noble Gases (b) O, Easily Available (c) F, High Electronegativity (d) Al, High Electropositivity

14. _____ is the maximum number of valence electrons a metal atom have (a) 4 (b) 3 (c) 2 (d) 1

15. A negative ion what is it sign? (a) Positive (b) Negative (c) Neutral (d) 2

16. _____ is the ability of a element to have same structural formula/molecular formula but different physical properties

17. Saturated in Organic Chemistry Means _____ (a) Partially Filled (b) Completely Filled (c) Selectively Filled (d) Irregularly Filled

18. Synthesis is Popularly Called (a) Complex (b) Simple (c) Backward Search (d) Soln

19. An _____ has same MASS number but different atomic numbers (a) Isomer (b) Isotopes (c) Allotropes (d) Isopropes

20. The Atomic theory was proposed by _____ (a) Prof. Sir Isaac Newton (b) Robert Hooke (c) Lord Rutherford (d) John Dalton

21. The Major Difference Between "Distillation" And Fractional Distillation Is (a) Wide Boiling Points of Liquids (b) Close boiling points (c) Solubility Curve (d) Range Of Boiling point

22. A Mixture May Be _____ (a) Molecular Formula (b) Structural Formula (c) Formula (d) Isotope

23. _____ Separating Technique involves "Mixture" (a) Magnetism (b) Distillation

(c) Polysilanes (d) Centrifugation

24. Which Of These is not a type of hybridization (a) sp^3 (b) sp^2 (c) sp (d) sp

25. What is The Valency OF Carbon (a) 4 (b) 3 (c) 5 (d) 2

26. Bonding Occurs between _____ of elements (a) Valence bond (b) Valence electrons (c) Valence Atom (d) Valence Molecules

27. _____ is saturated (a) C2H2 (b) C2H4 (c) C2H6 (d) C2H12

28. _____ And _____ Decomposes Becomes Water (a) Ethane And Methane (b) Ethane And Propane (c) Ethane And Pentane (d) Propane And Butane

29. _____ is the breakdown of a compound by the action of heat (a) Bonding (b) Cracking (c) Polymerization (d) Distillation

30. _____ is not a type of isomerism (a) Functional group (b) Chain (c) Molecular (d) Positional

31. There is a variation of _____ and _____ in the homologous (a) Physical and chemical properties (b) Molecular and properties (c) States and structures (d) Boiling and melting points

32. The Ability of an element to attract and hold water from the atmosphere is called _____ (a) Deliquescence (b) Coagulant (c) Hygroscopy (d) Desiccation

33. Destructive Distillation OF Coal is the heating of coal to very high temperatures in the absence of air to give _____ main products (a) 1 (b) 2 (c) 3 (d) 4

"I Wish You All The Best In Your Post-UTME Examination, I Pray Your Reaction Should Praise God And mercy In Jesus Name, Amen"

Chm 1tma Questions

Brendan G. Carr



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