

## Investigating Factors Affecting Enzyme Action **KEY**

### Teacher Notes:

In this lab, the assay system is a filter paper disc coated that is first coated with catalase and secondly dropped into a small container (vial) of hydrogen peroxide ( $\text{H}_2\text{O}_2$ ). The catalyst will break down the hydrogen peroxide into water and oxygen gas. As this happens, bubbles of oxygen will collect under the filter, causing it to rise to the top of the hydrogen peroxide. The rate of enzyme activity is known based on the time it takes for the filter to rise to the top.

**Preparation of Catalase Solution:** 50 g of peeled potato is mixed with 50 mL of cold distilled water and crushed ice; it is then homogenized in a blender for 30 seconds. This extract is then filtered through cheesecloth, and cold distilled water is added to bring it to a total volume of 100 mL. The enzyme should be kept on ice at all times. Another substance, besides potato (for example: liver), can also be used to extract catalase.

### Possible student setups:

**Enzyme Concentration:** The catalase solution can be diluted by the addition of cold distilled water.

**Substrate Concentration:** The hydrogen peroxide can be diluted by the addition of cold distilled water.

**Temperature:** Ice water, room temperature water, and warm and hot water baths can be used for different temperatures. Note that hydrogen peroxide should not be boiled, so refrain from using a boiling water bath.

**pH:** Different volumes of an acid and a base can be used to create a "highly acidic", "slightly acidic", "slightly alkaline", and "highly alkaline" pH environments.

### Analysis:

Answer the following questions in the analysis section of your laboratory report.

1. You measured the time it took for the catalase-soaked disk to float to the top. How does this amount of time indicate the amount of catalase activity happening in the solution?

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