



## Density Calculations Worksheet I

$$\text{density} = \frac{\text{mass}}{\text{volume}}$$

UNITS OF DENSITY  
solids (g/cm<sup>3</sup>)      liquids (g/mL)

1. Find the unknown quantity:

a)  $d = 3 \text{ g/mL}$   
 $V = 100 \text{ mL}$   
 $M = ?$   
 $m = d \times v$   
 $= (3 \text{ g/mL})(100 \text{ mL})$   
 $= 3 \times 10^2 \text{ g}$

b)  $d = ?$   
 $V = 950 \text{ mL}$   
 $M = 95 \text{ g}$   
 $d = \frac{m}{V}$   
 $= \frac{95 \text{ g}}{950 \text{ mL}} = 0.10 \text{ g/mL}$

c)  $d = 0.5 \text{ g/cm}^3$   
 $V = ?$   
 $M = 20 \text{ g}$   
 $V = \frac{m}{d}$   
 $= \frac{20 \text{ g}}{0.5 \text{ g/cm}^3} = 4 \times 10^1 \text{ cm}^3$

2. Find the unknown quantity (CONVERT FIRST to g or mL)

a)  $d = 24 \text{ g/mL}$   
 $V = 1.2 \text{ L} = \underline{1200} \text{ mL}$   
 $M = ?$   
 $m = d \times v$   
 $= (24 \text{ g/mL})(1200 \text{ mL})$   
 $= 2.9 \times 10^4 \text{ g}$

b)  $d = ?$   
 $V = 100 \text{ mL}$   
 $M = 1.5 \text{ kg} = \underline{1500} \text{ g}$   
 $d = \frac{m}{V}$   
 $= \frac{1500 \text{ g}}{100 \text{ mL}} = 15.0 \text{ g/mL}$

c)  $d = ?$   
 $V = 0.52 \text{ L} = \underline{520} \text{ mL}$   
 $M = 500 \text{ mg} = \underline{0.500} \text{ g}$   
 $d = \frac{m}{V}$   
 $= \frac{0.500 \text{ g}}{520 \text{ mL}} = 9.62 \times 10^{-4} \text{ g/mL}$

### WORD PROBLEMS

1. A block of aluminum occupies a volume of 15.0 mL and weighs 40.5 g. What is its density?

$V = 15.0 \text{ mL}$   
 $m = 40.5 \text{ g}$   
 $d = ?$   
 $d = \frac{m}{V}$   
 $= \frac{40.5 \text{ g}}{15.0 \text{ mL}}$

$d = 2.70 \text{ g/mL}$

2. Mercury metal is poured into a graduated cylinder that holds exactly 22.5 mL. The mercury used to fill the cylinder weighs 306.0 g. From this information, calculate the density of mercury.

$V = 22.5 \text{ mL}$   
 $m = 306.0 \text{ g}$   
 $d = ?$

$d = \frac{m}{V} = \frac{306.0 \text{ g}}{22.5 \text{ mL}} = 13.6 \text{ g/mL}$

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**Ensheng Dong**



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