

# Carrier System Design Manual

## Volume 1

Part 1 Load Estimating  
Part 2 Air Distribution  
Part 3 Piping Design



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# Carrier System Design Manual 12 Volume

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web using algebra the moles was founded to be 0.00352 moles of butane in order to find the molarity the mass of the butane should be divided by the moles after further calculations the molarity of the butane gas was calculated to be 56.81 grams moles this lab is done with tons of sources of error

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web this data shows that there was a gain the mass of the lighter which shows that there was a certain amount of butane 0.38 grams 0.02 grams gas that was successfully collected during the gas collection there is a mild influence of barometric



pressure during both trials as shown above in the raw data

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web 11 use avogadro s law to determine the number of moles of butane gas assume that butane is an ideal gas and that one mole has a volume of 22.4 l at stp v 1 n 1 v 2 n 2 0.0896 l n 22.4 l 1 mole n 0.0040 moles 12 experimental molar mass g mole of butane 0.23 g 0.0040 moles 58 g mole 13 the molecular formula of butane is  $C_4H_{10}$

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web nov 3 2020 in this experiment the moles and mass will be required to determine the molar mass of butane in a butane lighter experimental to begin the experiment the mass of a butane lighter was measured before using the butane gas to deplete the water to the 80ml mark on a graduated cylinder

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web dec 18 2020 in this lab students will experimentally determine the molar mass of a gas specifically butane  $C_4H_{10}$  by collection over water this experiment is an inquiry based experiment for 2nd year chemistry or ap chemistry students who have previously collected an insoluble gas

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web possible answer butane is extremely flammable and will ignite possible answer because the butane will not dissolve in the water it is possible for butane gas possible to form under water and get trapped possible answers 0.273 15 k and 1 atm 101.3 kpa  $C_4H_{10}$  12.01 g/mol 4.48 04 g/mol

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web the pressure of the gas the pressure of the gas was calculated using dalton s law of partial pressures by subtracting the pressure of the water vapor from the total pressure which in this case was the atmospheric pressure 760mm hg the pressure of the water vapor was found by using the temperature of the water

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web use the combined gas law to determine the volume in l of butane at stp use avogadro s law to determine the number of moles of butane gas assume that butane is an ideal gas and that one mole has a volume of 22 l at stp experimental molar mass g mole of butane the molecular formula of butane is  $C_4H_{10}$

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web although the most convenient gas for this is butane other gases may be available in similar small cylinders the collection of a gas sample and the weighing of the gas cylinder before and after this should take about 5 10 minutes as a demonstration

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