- What is the pH of a 0.0235 M HCl solution?
- What is the pOH of a 0.0235 M HCl solution?

What is the pH of a 6.50 x 10⁻³ M KOH solution? (Hint: this is a basic solution – concentration is of OH ')

4) A solution is created by measuring 3.60 x 10⁻³ moles of NaOH and 5.95 x 10⁻⁴ moles of HCl into a container and then water is added until the final volume is 1.00 L. What is the pH of this solution?

5) What is the pH of a 6.2 x 10⁻⁵ M NaOH solution? (Hint: this is a basic solution – concentration is of OH ')

6) A solution with a H* concentration of 1.00 x 10⁻⁷ M is said to be neutral. Why?

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Materials National Information Center for Special Education Materials, 1980

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