

Calculating pH

How is pH related to the concentration of hydronium ions?

Why?

In biology and other science courses pH is introduced as a way to quantify the acidity or basicity of a solution. This property can be measured using a pH probe or with an indicator paper strip that changes color at different pH values. But, what is actually being measured? We know that a pH of 7 is neutral, below 7 is acid, and above 7 is base, but why? What in the solution is the paper strip or probe actually reacting with?

Model 1 – Ion Concentrations for Acids and Bases

| Beaker | Solution | Acidic, Basic or Neutral? | $[H_3O^+]$ | $[OH^-]$ | $[H_3O^+] \times [OH^-]$ |
|--------|-----------------------|---------------------------|-------------------------|-------------------------|--------------------------|
| 1 | 0.10 M HCl(aq) | Acidic | 1.0×10^{-1} M | 1.0×10^{-13} M | 1×10^{-14} |
| 2 | 0.0010 M HCl(aq) | Acidic | 1.0×10^{-3} M | 1.0×10^{-11} M | 1×10^{-14} |
| 3 | 0.000010 M HCl(aq) | Acidic | 1.0×10^{-5} M | 1.0×10^{-9} M | 1×10^{-14} |
| 4 | 0.0000010 M HCl(aq) | Acidic | 1.0×10^{-6} M | 1.0×10^{-8} M | 1×10^{-14} |
| 5 | 0.00000010 M HCl(aq) | Neutral | 1.0×10^{-7} M | 1.0×10^{-7} M | 1×10^{-14} |
| 6 | 0.00000010 M NaOH(aq) | Neutral | 1.0×10^{-7} M | 1.0×10^{-7} M | 1×10^{-14} |
| 7 | 0.0000010 M NaOH(aq) | Basic | 1.0×10^{-8} M | 1.0×10^{-6} M | 1×10^{-14} |
| 8 | 0.00010 M NaOH(aq) | Basic | 1.0×10^{-10} M | 1.0×10^{-4} M | 1×10^{-14} |
| 9 | 0.010 M NaOH(aq) | Basic | 1.0×10^{-12} M | 1.0×10^{-2} M | 1×10^{-14} |
| 10 | 0.10 M NaOH(aq) | Basic | 1.0×10^{-13} M | 1.0×10^{-1} M | 1×10^{-14} |

1. What does the symbol $[H_3O^+]$ in Model 1 indicate?

The molar concentration of hydronium ion.

2. In Beaker 2, which ion has a higher concentration, hydronium ion or hydroxide ion?

Hydronium ion.

3. Describe how the concentration of hydronium ion was calculated for Beaker 3 in Model 1 from the concentration of the acid.

Since HCl is a strong acid, all of the molecules ionize in water. Therefore, the H_3O^+ concentration is equal to the initial acid concentration in the solution.

4. Describe how the concentration of hydroxide ion was calculated for Beaker 8 in Model 1 from the concentration of the base.

Since NaOH is a strong base, it completely dissociates in water. Therefore, the OH^- concentration is equal to the initial base concentration in the solution.

Calculating Ph Pogil Answer Key

Roger Gordon Bates



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A Generalised Equation for Calculating PH G. D. Peckham,1984 *Ph* Hans Skrifvars,1999 **Determination of PH** Roger Gordon Bates,1973 PH Measurements Victor Gold,1956 *Method of Determining PH by the Alkaline Absorption of Carbon Dioxide* ,1992 A method for measuring the concentration of hydroxides in alkaline solutions in a remote location using the tendency of hydroxides to absorb carbon dioxide The method includes the passing of carbon dioxide over the surface of an alkaline solution in a remote tank before and after measurements of the carbon dioxide solution A comparison of the measurements yields the absorption fraction from which the hydroxide concentration can be calculated using a correlation of hydroxide or pH to absorption fraction

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