

## Colligative Properties Worksheet – Answers

- 1) If I add 45 grams of sodium chloride to 500 grams of water, what will the melting and boiling points be of the resulting solution?  $K_b(\text{H}_2\text{O}) = 0.52^\circ\text{C}/\text{m}$  and  $K_f(\text{H}_2\text{O}) = 1.86^\circ\text{C}/\text{m}$ .

To find the melting and boiling points, you'll first need to find the concentration of the solution. To find the molality, convert grams of sodium chloride to moles (it turns out to be 0.769 moles) and divide by the number of kilograms of solvent (0.500 kg  $\text{H}_2\text{O}$ ). When you do this calculation, the molality is 1.54 m. Because sodium chloride breaks into two particles when it dissociates in water, the effective molality for the purposes of colligative property calculations is twice this, or 3.08 m.

From here, it's a matter of plugging the molality into the equations for freezing point depression and boiling point elevation. When you do this, you should get that the change in melting point is  $5.73^\circ\text{C}$  and the change in boiling point is  $1.60^\circ\text{C}$ .

The final answers: Melting point =  $-5.73^\circ\text{C}$ , boiling point =  $101.60^\circ\text{C}$ .

- 2) What is the vapor pressure of the solution in problem #1 at  $25^\circ\text{C}$ ? The vapor pressure of pure water at  $25^\circ\text{C}$  is 3.17 kPa.

Vapor pressure = mole fraction of water  $\times$  normal vapor pressure of water. The mole fraction of water in this case is equal to the number of moles of water (27.8) divided by the total number of moles of all substances in the solution ( $27.8 + 0.769$ ) = 28.6 moles. The mole fraction of water is 0.972. When you multiply the mole fraction times the normal vapor pressure, you find that the vapor pressure of the solution is 3.08 kPa.

- 3) Which solution will have a higher boiling point: A solution containing 105 grams of sucrose ( $\text{C}_{12}\text{H}_{22}\text{O}_{11}$ ) in 500 grams of water or a solution containing 35 grams of sodium chloride in 500 grams of water?

Using the same calculations as in question 1, you find that the boiling point of the sucrose solution is  $100.32^\circ\text{C}$  and the boiling point of the sodium chloride solution is  $101.40^\circ\text{C}$ . Clearly, the sodium chloride solution has a higher boiling point.

# Calculations Involving Colligative Properties Section Review Worksheet And Answers

**Robert C. Smoot, Smoot, Richard G.  
Smith, Jack Price**



**Calculations Involving Colligative Properties Section Review Worksheet And Answers:**

Merrill Chemistry Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998

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In some sort of consumed by monitors and the ceaseless chatter of immediate conversation, the melodic elegance and psychological symphony developed by the written term frequently diminish into the backdrop, eclipsed by the persistent sound and distractions that permeate our lives. Nevertheless, set within the pages of **Calculations Involving Colligative Propertiea Section Review Worksheet And Answers** an enchanting fictional prize full of fresh thoughts, lies an immersive symphony waiting to be embraced. Constructed by an outstanding composer of language, that captivating masterpiece conducts readers on an emotional journey, well unraveling the hidden songs and profound impact resonating within each carefully crafted phrase. Within the depths of this touching analysis, we will examine the book is key harmonies, analyze their enthralling publishing model, and submit ourselves to the profound resonance that echoes in the depths of readers souls.

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### **Table of Contents Calculations Involving Colligative Propertiea Section Review Worksheet And Answers**

1. Understanding the eBook Calculations Involving Colligative Propertiea Section Review Worksheet And Answers
  - The Rise of Digital Reading Calculations Involving Colligative Propertiea Section Review Worksheet And Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Calculations Involving Colligative Propertiea Section Review Worksheet And Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Calculations Involving Colligative Propertiea Section Review Worksheet And Answers

- User-Friendly Interface
- 4. Exploring eBook Recommendations from Calculations Involving Colligative Properties Section Review Worksheet And Answers
  - Personalized Recommendations
  - Calculations Involving Colligative Properties Section Review Worksheet And Answers User Reviews and Ratings
  - Calculations Involving Colligative Properties Section Review Worksheet And Answers and Bestseller Lists
- 5. Accessing Calculations Involving Colligative Properties Section Review Worksheet And Answers Free and Paid eBooks
  - Calculations Involving Colligative Properties Section Review Worksheet And Answers Public Domain eBooks
  - Calculations Involving Colligative Properties Section Review Worksheet And Answers eBook Subscription Services
  - Calculations Involving Colligative Properties Section Review Worksheet And Answers Budget-Friendly Options
- 6. Navigating Calculations Involving Colligative Properties Section Review Worksheet And Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Calculations Involving Colligative Properties Section Review Worksheet And Answers Compatibility with Devices
  - Calculations Involving Colligative Properties Section Review Worksheet And Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Calculations Involving Colligative Properties Section Review Worksheet And Answers
  - Highlighting and Note-Taking Calculations Involving Colligative Properties Section Review Worksheet And Answers
  - Interactive Elements Calculations Involving Colligative Properties Section Review Worksheet And Answers
- 8. Staying Engaged with Calculations Involving Colligative Properties Section Review Worksheet And Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Calculations Involving Colligative Properties Section Review Worksheet And Answers
- 9. Balancing eBooks and Physical Books Calculations Involving Colligative Properties Section Review Worksheet And Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Calculations Involving Colligative Properties Section Review Worksheet

And Answers

10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Calculations Involving Colligative Properties Section Review Worksheet And Answers
  - Setting Reading Goals Calculations Involving Colligative Properties Section Review Worksheet And Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Calculations Involving Colligative Properties Section Review Worksheet And Answers
  - Fact-Checking eBook Content of Calculations Involving Colligative Properties Section Review Worksheet And Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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