

Name \_\_\_\_\_

Section \_\_\_\_\_

**Experiment 3****Data and Calculations: Determination of the Equilibrium Constant for a Chemical Reaction**

Trials	Volume of 0.1M $\text{Fe}^{3+}$ (mL)	Volume of 0.1M $\text{SCN}^{-}$ (mL)	Volume of 0.1M $\text{H}^{+}$ (mL)	Initial $[\text{Fe}^{3+}]$ (M)	Initial $[\text{SCN}^{-}]$ (M)		Equilibrium $[\text{FeSCN}^{2+}]$ (M)
	Volume	Volume	Volume	Concentration	Concentration	Concentration	
1	5.00	1.00	2.00	0.010	0.002	0.000	0.000
2	5.00	1.00	3.00	0.010	0.002	0.000	0.000
3	5.00	1.00	2.00	0.010	0.002	0.000	0.000
4	5.00	1.00	1.00	0.010	0.002	0.000	0.000
5	5.00	1.00	0.00	0.010	0.002	0.000	0.000

If Method 1 is used, Section 1 of the report should be completed. If Method 2 is used, Section 2 of the report should be completed.

**Processing the Data**

A. Calculation of  $K_c$  assuming the reaction:



This calculation is more easily done by knowing how to change  $K_c$  to the reaction  $\text{FeSCN}^{2+} \rightleftharpoons \text{Fe}^{3+} + \text{SCN}^{-}$  if you are using Method 1 or the reverse of the above, and follow the directions that follow.

**Step 1:** Find the initial number of moles of  $\text{Fe}^{3+}$  and  $\text{SCN}^{-}$  in the mixture (in mL) used (Section 1) or (Section 2) for each mixture in the three mixtures in the table.

**Step 2:** From the experimentally determined value of  $[\text{FeSCN}^{2+}]$  at equilibrium for each of the mixtures, it follows by stoichiometry that the initial number of moles of  $\text{Fe}^{3+}$  and  $\text{SCN}^{-}$  in each of the mixtures is the same as the initial number of moles of  $\text{FeSCN}^{2+}$  at equilibrium. This is the case for the reaction  $\text{Fe}^{3+} + \text{SCN}^{-} \rightleftharpoons \text{FeSCN}^{2+}$ .

continued on following page

# Calculations With A Chemical Reaction Lab Answers

**JL Elias**



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