

Specific Heat Equation

$$Q = mC\Delta T$$

Q = thermal energy

m = mass

C = specific heat

ΔT = change in temp

Ex: The specific heat of silicon is $703 \text{ J}/(\text{kg} \cdot ^\circ\text{C})$. How much energy is needed to raise a 7 kg chunk of silicon 10°C ?

answer: $Q = 7 \text{ kg} \cdot \frac{703 \text{ J}}{\text{kg} \cdot ^\circ\text{C}} \cdot 10^\circ\text{C} = 49210 \text{ J}$

Note that the units do indeed work out to be energy units.

**Use table 12-1 on p. 248 for some specific heats of different substances.*

Calculating Specific Heat Answers

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