

**Chemistry Unit 15: Chemical Equilibrium Worksheet****Practice 15.1 Writing Equilibrium Equations**

Write the equilibrium expressions ( $K_{eq}$ ) for the following reactions

1.  $2\text{CO(g)} \leftrightarrow \text{C(s)} + \text{CO}_2\text{(g)}$
2.  $\text{N}_2\text{O}_4\text{(g)} \leftrightarrow 2\text{NO}_2\text{(g)}$
3.  $2\text{HCl(g)} \leftrightarrow \text{I}_2\text{(g)} + \text{Cl}_2\text{(g)}$
4.  $\text{Fe}_2\text{O}_3\text{(s)} + 3\text{H}_2\text{(g)} \leftrightarrow 2\text{Fe(s)} + 3\text{H}_2\text{O(g)}$
5.  $2\text{CO(g)} + 4\text{H}_2\text{O(g)} \leftrightarrow 3\text{O}_2\text{(g)} + 2\text{CH}_4\text{(g)}$
6.  $\text{P}_4\text{(s)} + 6\text{NO(g)} \leftrightarrow \text{P}_4\text{O}_6\text{(s)} + 3\text{N}_2\text{(g)}$

**15.2 Calculating Equilibrium Constants**

Be sure to write out the equilibrium ( $K_{eq}$ ) expression for each problem before trying to solve it!

1. What is the equilibrium expression for the following reaction?  
 $\text{SeO}_2\text{(s)} + 2\text{CO(g)} \leftrightarrow \text{Se(s)} + 2\text{CO}_2\text{(g)}$
2. At equilibrium the following concentrations are recorded: 20.0 M of  $\text{H}_2$ , 18.0 M of  $\text{CO}_2$ , 12.0 M  $\text{H}_2\text{O}$ , and 5.9 M of  $\text{CO}$  at 427°C. What is the value of  $K_{eq}$  at this temperature for the following reaction?  
 $\text{CO}_2\text{(g)} + \text{H}_2\text{(g)} \leftrightarrow \text{CO(g)} + \text{H}_2\text{O(g)}$
3. At equilibrium a mixture of  $\text{N}_2$ ,  $\text{H}_2$ , and  $\text{NH}_3$  gas at 500°C is determined to consist of 0.602 M  $\text{N}_2$ , 0.420 M  $\text{H}_2$ , and 0.113 M of  $\text{NH}_3$ . What is the equilibrium constant at this temperature? Does this equilibrium favor the reactants or the products?  
 $\text{N}_2\text{(g)} + 3\text{H}_2\text{(g)} \leftrightarrow 2\text{NH}_3\text{(g)}$
4. Given the equilibrium concentrations of 0.15 M  $\text{NO}$ , 0.12 M  $\text{H}_2$  and 1.6 M  $\text{N}_2$ . Calculate the equilibrium constant for the following reaction. Does this reaction favor the reactants or the products?  
 $2\text{NO(g)} + 2\text{H}_2\text{(g)} \leftrightarrow \text{N}_2\text{(g)} + 2\text{H}_2\text{O(l)}$
5. What is the equilibrium constant for the following reaction if the equilibrium concentrations for the gases are 0.3 M for  $\text{C}_2\text{H}_6$ , 0.6 M for  $\text{O}_2$ , 1.3 M for  $\text{CO}_2$  and 1.8 M for  $\text{H}_2\text{O}$ ? Does the reaction favor the reactants or products?  
 $2\text{C}_2\text{H}_6\text{(g)} + 7\text{O}_2\text{(g)} \leftrightarrow 4\text{CO}_2\text{(g)} + 6\text{H}_2\text{O(g)}$
6. Nitrogen dioxide reacts to form dinitrogen tetraoxide. What is the equilibrium concentration of  $\text{N}_2\text{O}_4$  if the  $K_{eq} = 4.70$  and the equilibrium concentration of  $\text{NO}_2$  is 0.0310 M?  
 $2\text{NO}_2\text{(g)} \leftrightarrow \text{N}_2\text{O}_4\text{(g)}$

# Calculating Equilibrium Constants Key Chem Worksheet 18 3

**Brian Coleman**



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