



Name: \_\_\_\_\_  
Date: \_\_\_\_\_

# COUNTING ATOMS

## WORKSHEET

- Chemical formulas can have three components.
  - The \_\_\_\_\_ which are represented by a single uppercase letter, or represented by an uppercase letter followed by a lowercase letter. This tells you the types of elements in the compound.
  - The \_\_\_\_\_ which are numbers that are found on the lower \_\_\_\_\_ - hand side of each element symbol. This tells you the number of \_\_\_\_\_ of this element in the molecule. If an element symbol has no subscript next to it, then this indicates that there is only \_\_\_\_\_ atom of this element in the molecule.
  - The \_\_\_\_\_ which surround some groups of atoms indicate that the numbers of all of the atoms inside the bracket need to be \_\_\_\_\_ by the \_\_\_\_\_ on the outside of the bracket found on the lower \_\_\_\_\_ - hand side of that bracket.
- Record the number of each atom in each molecule, then record the total number of atoms in the molecule:

| 1) NaOH |            |
|---------|------------|
| Element | # of Atoms |
|         |            |
|         |            |
| Total   |            |

| 4) Li <sub>2</sub> SO <sub>4</sub> |            |
|------------------------------------|------------|
| Element                            | # of Atoms |
|                                    |            |
|                                    |            |
| Total                              |            |

| 7) H <sub>2</sub> COCH <sub>2</sub> |            |
|-------------------------------------|------------|
| Element                             | # of Atoms |
|                                     |            |
|                                     |            |
| Total                               |            |

| 10) NH <sub>4</sub> C <sub>2</sub> H <sub>3</sub> O <sub>2</sub> |            |
|--|------------|
| Element  | # of Atoms |
|  |            |
|  |            |
| Total  |            |

| 13) Ca <sub>3</sub> (PO <sub>4</sub> ) <sub>2</sub> |            |
|---|------------|
| Element   | # of Atoms |
|   |            |
|   |            |
| Total   |            |

| 2) HNO <sub>3</sub> |            |
|---------------------|------------|
| Element             | # of Atoms |
|                     |            |
|                     |            |
| Total               |            |

| 5) NH <sub>4</sub> Cl |            |
|-----------------------|------------|
| Element               | # of Atoms |
|                       |            |
|                       |            |
| Total                 |            |

| 8) Mg(OH) <sub>2</sub> |            |
|------------------------|------------|
| Element                | # of Atoms |
|                        |            |
|                        |            |
| Total                  |            |

| 11) NaC <sub>2</sub> H <sub>3</sub> O <sub>2</sub> |            |
|--|------------|
| Element  | # of Atoms |
|  |            |
|  |            |
| Total  |            |

| 14) Fe <sub>2</sub> (CO <sub>3</sub> ) <sub>3</sub> |            |
|---|------------|
| Element   | # of Atoms |
|   |            |
|   |            |
| Total   |            |

| 3) U <sub>2</sub> O |            |
|---------------------|------------|
| Element             | # of Atoms |
|                     |            |
|                     |            |
| Total               |            |

| 6) CaClO <sub>3</sub> |            |
|-----------------------|------------|
| Element               | # of Atoms |
|                       |            |
|                       |            |
| Total                 |            |

| 9) Al(OH) <sub>3</sub> |            |
|------------------------|------------|
| Element                | # of Atoms |
|                        |            |
|                        |            |
| Total                  |            |

| 12) (NH <sub>4</sub> ) <sub>3</sub> PO <sub>4</sub> |            |
|---|------------|
| Element   | # of Atoms |
|   |            |
|   |            |
| Total   |            |

| 15) Al <sub>2</sub> (SO <sub>4</sub> ) <sub>3</sub> |            |
|---|------------|
| Element   | # of Atoms |
|   |            |
|   |            |
| Total   |            |

# Counting Atoms Study Guide Answers

**Karen Timberlake**



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creative problems This edition contains over 1 800 problems many of them new and taken directly from the scientific literature The book is used at a wide variety of schools such as UC Berkeley Caltech Colorado Cornell Duke Harvard Illinois Maryland Purdue Yale Wisconsin and many more This edition provides students with more health examples drawn from modern medical practice as well as many cutting edge topics from modern synthetic organic chemistry In addition to the printed book students can rely on Sapling Learning's online homework platform for extra learning and assessment The platform offers automatic grading an easy to use interface and instructive feedback Instructors can select from a variety of existing problem sets over 1 000 of Loudon's problems are in the platform or they can modify the questions or author them from scratch Not only does the software allow students to easily draw and interact with structures it allows them to draw entire reaction mechanisms including showing the movement of electrons with curved electron arrows

Student Study Guide to accompany Foundations of College Chemistry, 11th Edition Morris Hein, Susan Arena, 2003-09-03 Work more effectively and gauge your progress along the way This Student Study Guide that accompanies Hein's Foundations of College Chemistry 11th Edition contains a review and self evaluation section challenge problems and answers and solutions for all exercises in the study guide Foundations of College Chemistry is the book that defined the prep intro chemistry market 35 years ago And it's been a market leader ever since Hein Arena is known for its accuracy clear no nonsense approach and direct writing style Strong problem solving and carefully constructed problem sets make this book a standout among its many imitators

**Study Guide for Hein and Arena's Foundations of College Chemistry** Peter Scott, Rachael Enriquez Porter, 1999-04-18

*Study Guide & Solutions Manual to Accompany Organic Chemistry* Mary H. Bailey, 1998

Student Study Guide Fundamentals of Chemistry by David E. Goldberg Margaret G. Kimble, 1994

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