

CHAPTER 6 REVIEW*Chemical Bonding***MIXED REVIEW**

SHORT ANSWER Answer the following questions in the space provided.

1. Name the type of energy that is a measure of strength for each of the following types of bonds:

_____ lattice energy _____	a. ionic bond
_____ bond energy _____	b. covalent bond
_____ enthalpy of vaporization _____	c. metallic bond

2. Use the electronegativity values shown in **Figure 20**, on page 161 of the text, to determine whether each of the following bonds is nonpolar covalent, polar covalent, or ionic.

_____ ionic _____	a. H—F	_____ nonpolar covalent _____	d. H—H
_____ ionic _____	b. Na—Cl	_____ polar covalent _____	e. H—C
_____ polar covalent _____	c. H—O	_____ polar covalent _____	f. H—N

3. How is a hydrogen bond different from an ionic or covalent bond?

A hydrogen bond is a dipole-dipole attraction between a partially positive hydrogen
atom and the unshared electron pair of a strongly electronegative atom such as O,
N, or F. Unlike ionic or covalent bonds, in which electrons are given up or shared,
the hydrogen bond is a weaker attraction. Hydrogen bonds are generally
intermolecular, while ionic and covalent bonds occur between ions or atoms
respectively.

4. H₂S and H₂O have similar structures and their central atoms belong to the same group. Yet H₂S is a gas at room temperature and H₂O is a liquid. Use bonding principles to explain why this is.

Oxygen has higher electronegativity than sulfur, which creates a highly polar bond.
Increased polarity in H₂O bonds means a stronger intermolecular attraction, making
water a liquid at room temperature. Hydrogen bonding exists between water
molecules, but not between hydrogen sulfide molecules.

Covalent Bonding Worksheet Answers Chapter 8

S Marginson



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Chemical Misconceptions Keith Taber, 2002 Part 2 provides strategies for dealing with some of the misconceptions that students have by including ready to use classroom resources **Books in Print Supplement** ,2002 **Educart ICSE Class 10 One-shot Question Bank 2026 Chemistry (strictly for 2025-26 boards)** Sir Tarun Rupani, 2025-07-12 Fast track your Chemistry revision with this exam ready resource This One shot Question Bank by Sir Tarun Rupani is designed to help ICSE Class 10 students revise the complete Chemistry syllabus quickly and thoroughly It simplifies theory boosts numerical accuracy and ensures strong exam practice all aligned with the 2025 26 ICSE syllabus Key Features Strictly Based on ICSE 2025 26 Curriculum Complete chapter coverage including Periodic Table Chemical Bonding Acid Base Organic Chemistry and more One shot Format Each chapter includes concise concept notes chemical equations reactions and key diagrams for quick recall Complete Coverage of Question Types Includes objective short long answers equation based numerical and reasoning questions Chapterwise PYQs Included Practice with previous years ICSE board questions to understand trends and improve retention Solved Answers in ICSE Format Clear well structured solutions using proper units chemical symbols and balanced equations Smart Revision Focus Special tips to avoid common mistakes in writing reactions balancing equations and attempting numericals Why Choose This Book This Chemistry One shot by Sir Tarun Rupani is built for smart preparation whether you re revising at the last minute or practising throughout the term It helps you approach each question with clarity confidence and the precision needed to score high in the 2026 ICSE board exam *Prentice Hall Physical Science Concepts in Action Program Planner National Chemistry Physics Earth Science* ,2003-11 Prentice Hall Physical Science Concepts in Action helps students make the important connection between the science they read and what they experience every day Relevant content lively explorations and a wealth of hands on activities take students understanding of science beyond the page and into the world around them Now includes even more technology tools and activities to support differentiated instruction Exploring Earth and Space Michael DiSpezio, 1995 A textbook exploring such aspects of matter and energy as heat electricity and nuclear chemistry with suggested activities and review questions at the end of each chapter **Addison-Wesley Science Insights** ,1996 *Science Spectrum* Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2003-03 MnM_POW-Science-PM-10 (Updated) Vibha Arora, Anju Sachdeva, Sushma Sardana, MnM_POW Science PM 10 Updated **Earth Science** Samuel N. Namowitz, Nancy E. Spaulding, 1989 **Chapter Resource 2 Chemistry of Life Biology** Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2004 *Discover Science: Teacher's annotated edition* ,1991 Science content helps develop the skills needed to understand how science works learn new concepts solve problems and make decisions in today s technological society Holt Science and Technology Holt, Rinehart and Winston Staff, 2001 **Chemical Bonding Clarified Through Quantum Mechanics** George C. Pimentel, Richard D. Spratley, 1969

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