

Assigning Oxidation Numbers

Rule	Example
1. The oxidation number of any uncombined element is 0.	The oxidation number of $\text{Na}(s)$ is 0.
2. The oxidation number of a monatomic ion equals the charge on the ion.	The oxidation number of Cl^- is -1.
3. The more-electronegative element in a binary compound is assigned the number equal to the charge it would have if it were an ion.	The oxidation number of O in NO is -2.
4. The oxidation number of fluorine in a compound is always -1.	The oxidation number of F in LiF is -1.
5. Oxygen has an oxidation number of -2 unless it is combined with F, in which it is +1 or +2, or it is in a peroxide, in which it is -1.	The oxidation number of O in NO_2 is -2.
6. Hydrogen's oxidation state in most of its compounds is +1 unless it is combined with a metal, in which case it is -1.	The oxidation number of H in LiH is -1.
7. In compounds, the elements of Groups 1 and 2 as well as aluminum have oxidation numbers of +1, +2, and +3, respectively.	The oxidation number of Ca in CaCO_3 is +2.
8. The sum of the oxidation numbers of all atoms in a neutral compound is 0.	The oxidation number of C in CaCO_3 is +4.
9. The sum of the oxidation numbers of all atoms in a polyatomic ion equals the charge of the ion.	The oxidation number of P in H_2PO_4^- is +5.

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Guillaume Favre



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Physical Chemistry Peter Atkins, Julio de Paula, 2002 New edition of the overwhelmingly favorite text for the physical chemistry course **Principles and Practice of Ophthalmology E-Book** Daniel M. Albert, Joan W. Miller, Dimitri T. Azar, Barbara A. Blodi, 2008-02-27 Inside the 3rd edition of this esteemed masterwork hundreds of the most distinguished authorities from around the world provide today's best answers to every question that arises in your practice They deliver in depth guidance on new diagnostic approaches operative technique and treatment option as well as cogent explanations of every new scientific concept and its clinical importance With its new streamlined more user friendly full color format this 3rd edition makes reference much faster easier and more versatile More than ever it is the source you need to efficiently and confidently overcome any clinical challenge you may face Comprehensive authoritative and richly illustrated coverage of every scientific and clinical principle in ophthalmology ensures that you will always be able to find the guidance you need to diagnose and manage your patients ocular problems and meet today's standards of care Updates include completely new sections on Refractive Surgery and Ethics and Professionalism an updated and expanded Genetics section an updated Retina section featuring OCT imaging and new drug therapies for macular degeneration and many other important new developments that affect your patient care A streamlined format and a new more user friendly full color design with many at a glance summary tables algorithms boxes diagrams and thousands of phenomenal color illustrations allows you to locate the assistance you need more rapidly than ever **Fundamentals of Chemistry** Ralph A. Burns, 2003 For one semester preparatory chemistry courses or general purpose introductory chemistry courses This clearly written well illustrated versatile textbook provides thorough coverage of chemistry with a balance of

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