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Determining Empirical Formulas Answer: 75% carbon, 25% hydrogen CH4 Work: C 751. > 75gC | 1 mol = 6.24 mol/6.24 = 1 | C Hq 2) 52.7% potassium, 47.3% chlorine Answer: KCL R 5271 - 52.79K | 1 mol K = 1.34 mol/133 = 1 CI 47.37 > 47.39 CI I I MOI CI = 1.33 MOI /1.83 = 1 22.1% aluminum, 25.4% phosphorous, 52.5 % oxygen Answer: AIPO+ A1 22.17. -> 22.19 A1 1 mol A1 - 0.819 mol/0.819 = 1 AIPO. P 2547. -> 25.49P | 1 molP = 0.820 mol/0.819 = 1 0 5251 = 52590 | 1mo10 = 3.28 mo1/0.819 = 4 4) 13% magnesium, 87% bromine | A Answer: MaBra Work: 13 g mg | 1 mol mg = 0.5 5 mol mg MaBrz 879Br | 1moi Br = 1.09/054 = Z NazSD+ 32.4% sodium, 22.5% sulfur, 45.1% oxygen Answer: 324gNa | 1mol = 1.41mol Nayo.70 = 2 Naz SO4 Work 22.595 | 1mols = 0.70 mols/070= 1 45.190 | 1mo10 = 2.82mo10/0.70 = 4

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**Léo Pomar** 

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