

DETERMINING MOLECULAR FORMULAS (TRUE FORMULAS)

Name: _____

Derive the molecular formula:

1. The empirical formula of a compound is PCl_2 . Its molecular mass is 263 g/mol. What is its molecular formula?

Given: Empirical formula =

Mass

2. The empirical formula of a compound is CH_2 . Its molecular mass is 78 g/mol. What is its molecular formula?

Given: Empirical formula = Mass of empirical formula =

Mass

3. A compound is found to contain 40.0% carbon, 5.15% hydrogen and 54.8% oxygen. Its molecular mass is 176 g/mol. What is its molecular formula?

Given: % Carbon = % Hydrogen = % Oxygen =

Given: Molar mass =

Empirical formula =

Mass

4. A compound is found to contain 54.5% carbon, 3.75% hydrogen and 41.7% oxygen. Its molecular mass is 176 g/mol. What is its molecular formula?

Given: % Carbon = % Hydrogen = % Oxygen =

Given: Molar mass =

Empirical formula =

Mass

5. A compound is found to contain 68.8% carbon, 11.9% hydrogen and 19.3% oxygen. Its molecular mass is 176 g/mol. What is its molecular formula?

Given: % Carbon = % Hydrogen = % Oxygen =

Given: Molar mass =

Empirical formula =

Mass

Determining Molecular Formulas True Formulas Answer Key

D Kirk



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