Section 6.1: Organizing the Elements

- 1. Group I on the periodic table is called Alkali Metals
- 2. Group 2 on the periodic table is called Alkaline Earth Metals
- 3. Elements in groups 3-12 are called Transition Metals
- Group 17 on the periodic table is called Halogens
- 5. Group 18 on the periodic table is called Noble Gases
- Each period on the periodic table corresponds to a principal energy level
- 7. The 2 series that make up the inner transition metals are Lanthanides and Actinides
- Complete the following chart:

Element	Symbol	Group #	Period #	Group Name	Metal, Metalloid, Nonmetal
Zirconium	Zr	- 4	5	Transition Metals	Metal
Radium	Ra	2	7	Alkaline Earth Metals	Metal
Bismuth	Bi	1.5	6	Nitrogen Group	Metal
Gallium	Ga	1.3	4	Boron Group	Metal

- 9. What are triads? Sets of 3 elements with similar properties Whose idea were they? Dobereiner
- 10. The modern periodic table is arranged by increasing atomic number. Whose idea was this? Moseley
- Mendeleev developed the first periodic table, which was arrange by increasing atomic mass
- 12. Determine which of the following pairs will have similar properties (Yes or no):
 - a. Ca and Be? Yes same group (Group 2 Alkaline Earth Metals)
 - b. Sr and Y? No not the same group
 - c. C and Ge? Yes same group (Group 14 Carbon Group)

Section 6.3: Periodic Trends

- As you go across a period, the atomic radius (size) decreases because the attraction between protons
 and electrons increases, which makes the atom smaller (opposite of tug of war)
- As you go down a group, the atomic radius (size) increases because the number of principal energy levels increases
- Circle the atom with the largest atomic radius and put a square around the atom with the smallest radius:
 - a. (Li) Ne F B
 b. N Bi S P
 c. Na (Sr) Al (Cl
- 4. Ionization energy is the energy required to remove an electron.

Chapter 6 1 Organizing The Elements Worksheet Answers

Manish Soni

Chapter 6 1 Organizing The Elements Worksheet Answers:

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Table of Contents Chapter 6 1 Organizing The Elements Worksheet Answers

- 1. Understanding the eBook Chapter 6 1 Organizing The Elements Worksheet Answers
 - The Rise of Digital Reading Chapter 6 1 Organizing The Elements Worksheet Answers
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Chapter 6 1 Organizing The Elements Worksheet Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 6 1 Organizing The Elements Worksheet Answers
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Chapter 6 1 Organizing The Elements Worksheet Answers
 - Personalized Recommendations
 - Chapter 6 1 Organizing The Elements Worksheet Answers User Reviews and Ratings

- Chapter 6 1 Organizing The Elements Worksheet Answers and Bestseller Lists
- 5. Accessing Chapter 6 1 Organizing The Elements Worksheet Answers Free and Paid eBooks
 - Chapter 6 1 Organizing The Elements Worksheet Answers Public Domain eBooks
 - Chapter 6 1 Organizing The Elements Worksheet Answers eBook Subscription Services
 - Chapter 6 1 Organizing The Elements Worksheet Answers Budget-Friendly Options
- 6. Navigating Chapter 6 1 Organizing The Elements Worksheet Answers eBook Formats
 - o ePub, PDF, MOBI, and More
 - Chapter 6 1 Organizing The Elements Worksheet Answers Compatibility with Devices
 - Chapter 6 1 Organizing The Elements Worksheet Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 6 1 Organizing The Elements Worksheet Answers
 - Highlighting and Note-Taking Chapter 6 1 Organizing The Elements Worksheet Answers
 - Interactive Elements Chapter 6 1 Organizing The Elements Worksheet Answers
- 8. Staying Engaged with Chapter 6 1 Organizing The Elements Worksheet Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 6 1 Organizing The Elements Worksheet Answers
- 9. Balancing eBooks and Physical Books Chapter 6 1 Organizing The Elements Worksheet Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 6 1 Organizing The Elements Worksheet Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chapter 6 1 Organizing The Elements Worksheet Answers
 - Setting Reading Goals Chapter 6 1 Organizing The Elements Worksheet Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chapter 6 1 Organizing The Elements Worksheet Answers
 - Fact-Checking eBook Content of Chapter 6 1 Organizing The Elements Worksheet Answers
 - Distinguishing Credible Sources

- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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