

Redox practice worksheet

Name: _____

Date: _____

- In which substance is the oxidation number of nitrogen zero?
A. NH_3 B. N_2 C. NO_2 D. N_2O
- What is the oxidation number of carbon in NaHCO_3 ?
A. +6 B. +2 C. -4 D. +4
- In the reaction $\text{Al}^0 + \text{Cr}^{3+} \rightarrow \text{Al}^{3+} + \text{Cr}^0$, the reducing agent is
A. Al^0 B. Cr^{3+} C. Al^{3+} D. Cr^0

- In the reaction $2\text{K} + \text{Cl}_2 \rightarrow 2\text{KCl}$, the species oxidized is
A. Cl_2 B. Cl^- C. K D. K^+
- As an S^{2-} ion is oxidized to an S^0 atom, the number of protons in its nucleus
A. decreases B. increases
C. remains the same

- Given the probable reaction for the nickel-cadmium battery:
$$2\text{Ni}(\text{OH})_2 + \text{Cd} \rightleftharpoons 2\text{Ni}(\text{OH})_2 + \text{Cd}(\text{OH})_2$$

Which species is oxidized during the discharge of the battery?

A. Ni^{2+} B. Ni^{2+} C. Cd^0 D. Cd^{2+}

- If element X forms the oxides XO and X_2O_3 , the oxidation numbers of element X are
A. +1 and +2 B. +2 and +3
C. +1 and +3 D. +2 and +4
- Oxygen has a positive oxidation number in the compound
A. H_2O B. H_2O_2 C. OF_2 D. IO_2
- What is the oxidation number of sulfur in H_2SO_4 ?
A. 0 B. -2 C. +6 D. +4

- In the equation $\text{Cu(s)} + 2\text{Ag}^+(\text{aq}) \rightarrow \text{Cu}^{2+}(\text{aq}) + 2\text{Ag(s)}$, the oxidizing agent is
A. Cu^0 B. Ag^+ C. Cu^{2+} D. Ag^0
- In the reaction $\text{Al} + \text{Cr}^{3+} \rightarrow \text{Al}^{3+} + \text{Cr}$, the reducing agent is
A. Al B. Cr^{3+} C. Al^{3+} D. Cr
- In the compound Na_2HPO_4 , which element has a negative oxidation number?
A. H B. O C. P D. Na

Chapter 2 worksheet Redox Answers

Yan Bai

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the name 'Yan Bai'.

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Table of Contents Chapter 2worksheet Redox Answers

1. Understanding the eBook Chapter 2worksheet Redox Answers
 - The Rise of Digital Reading Chapter 2worksheet Redox Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 2worksheet Redox Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 2worksheet Redox Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 2worksheet Redox Answers
 - Personalized Recommendations
 - Chapter 2worksheet Redox Answers User Reviews and Ratings
 - Chapter 2worksheet Redox Answers and Bestseller Lists

5. Accessing Chapter 2 worksheet Redox Answers Free and Paid eBooks
 - Chapter 2 worksheet Redox Answers Public Domain eBooks
 - Chapter 2 worksheet Redox Answers eBook Subscription Services
 - Chapter 2 worksheet Redox Answers Budget-Friendly Options
6. Navigating Chapter 2 worksheet Redox Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 2 worksheet Redox Answers Compatibility with Devices
 - Chapter 2 worksheet Redox Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 2 worksheet Redox Answers
 - Highlighting and Note-Taking Chapter 2 worksheet Redox Answers
 - Interactive Elements Chapter 2 worksheet Redox Answers
8. Staying Engaged with Chapter 2 worksheet Redox Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 2 worksheet Redox Answers
9. Balancing eBooks and Physical Books Chapter 2 worksheet Redox Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 2 worksheet Redox Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chapter 2 worksheet Redox Answers
 - Setting Reading Goals Chapter 2 worksheet Redox Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chapter 2 worksheet Redox Answers
 - Fact-Checking eBook Content of Chapter 2 worksheet Redox Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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